Abstract: The development of technetium chemistry has been lagging behind that of its heavier congener rhenium, primarily because the inherent radioactivity of all Tc isotopes has limited the number of laboratories that can study the chemistry of this fascinating element. Although technetium is an artificial element, it is not rare. Significant amounts of the isotope (99)Tc are produced every day as a fission byproduct in nuclear power plants. Therefore, a fundamental understanding of the chemistry of (99)Tc is essential to avoid its release into the environment. In this article the chemistry of technetium at its highest oxidation state (+VII) is reviewed with a special focus on recent developments which make water-stable complexes of the general type TcO(3)(tacn-R)(+) (tacn-R = 1,4,7-triazacyclononane or derivatives) accessible. Complexes containing the fac-TcO(3)(+) core display a unique reactivity. In analogy to OsO(4) and RuO(4), complexes containing the fac-TcO(3)(+) core undergo with alkenes metal-mediated, vicinal cis-dihydroxylation reactions (alkene-glycol interconversion) in water via a (3+2)-cycloaddition reaction. Therefore, water-stable fac-(99m)TcO(3)(+) complexes pave the way for a new labeling strategy for radiopharmaceutical applications, based on (3+2)-cycloaddition reactions. This new concept for the labeling of biomolecules with small (99m)TcO(3)(tacn-R)(+)-type complexes by way of a (3+2)-cycloaddition with alkenes is discussed in detail. The herein reported developments in high-valent technetium chemistry create a new field of research with this artificial element. This demonstrates the potential of fundamental research to provide new impetus of innovation for the development of new methods for radiopharmaceutical applications.

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Water-stable fac-\{\text{TcO}_3\}^+ Complexes – A New Field of Technetium Chemistry

Henrik Braband*

Abstract: The development of technetium chemistry has been lagging behind that of its heavier congener rhenium, primarily because the inherent radioactivity of all Tc isotopes has limited the number of laboratories that can study the chemistry of this fascinating element. Although technetium is an artificial element, it is not rare. Significant amounts of the isotope $^{\text{99}}\text{Tc}$ are produced every day as a fission byproduct in nuclear power plants. Therefore, a fundamental understanding of the chemistry of $^{\text{99}}\text{Tc}$ is essential to avoid its release into the environment. In this article the chemistry of technetium at its highest oxidation state (+VII) is reviewed with a special focus on recent developments which make water-stable complexes of the general type $[\text{TcO}_3(tacn-R)]^+$ (tacn-R = 1,4,7-triazacyclononane or derivatives) accessible. Complexes containing the fac-\{\text{TcO}_3\}^+ core display a unique reactivity. In analogy to $[\text{OsO}_3]$ and $[\text{RuO}_3]$, complexes containing the fac-\{\text{TcO}_3\}^+ core undergo with alkenes metal-mediated, vicinal cis-dihydroxylation reactions (alkene-glycol interconversion) in water via a (3+2)-cycloaddition reaction. Therefore, water-stable fac-\{\text{TcO}_3\}^+ complexes pave the way for a new labeling strategy for radiopharmaceutical applications, based on (3+2)-cycloaddition reactions. This new concept for the labeling of biomolecules with small \{\text{TcO}_3(tacn-R)\}^+-type complexes by way of a (3+2)-cycloaddition with alkenes is discussed in detail. The herein reported developments in high-valent technetium chemistry create a new field of research with this artificial element. This demonstrates the potential of fundamental research to provide new impetus of innovation for the development of new methods for radiopharmaceutical applications.

Keywords: Alkenes · (3+2)-Cycloaddition · Labeling · Radiopharmacy · Technetium

1. Introduction

Technetium is a fascinating element, located in the middle of the periodic table (element 43). Apart from the fact that it does not have even one stable isotope, it can be characterized as a typical second row transition metal. Technetium is not a member of the natural decay series of uranium and thorium, therefore, it is often referred to as a ‘man-made’ element. Although technetium is an artificial element, it is not rare. Technetium is a fission byproduct in nuclear power plants. It was estimated that a typical pressurized water reactor (PWR), which is one of two types of light water reactors, produces about 9.3 kg of the isotope technetium-99 ($^{\text{99}}\text{Tc}$) per gigawatt year.

Based on this estimation and the amount of energy supplied by nuclear power plants worldwide, it can be calculated that today approximately 9.5 kg of $^{\text{99}}\text{Tc}$ are produced every day. $^{\text{99}}\text{Tc}$ is separated from used nuclear fuel in a reprocessing method known as the Plutonium-Uranium Extraction process (PUREX process). This leads to a current $^{\text{99}}\text{Tc}$ ‘production’ of over three tonnes per year. Fig. 1 shows the estimated $^{\text{99}}\text{Tc}$ production worldwide from 1970 until 2009 based on the capacity of all nuclear power plants. Furthermore, it shows the possible development of $^{\text{99}}\text{Tc}$ production (low and high estimate) until 2050 based on calculations for needs of energy, electricity and nuclear power made by the International Atomic Energy Agency (IAEA). This should demonstrate the abundance of this radioactive artificial element today and in the near future.

$^{\text{99}}\text{Tc}$ is a weak $\beta$-emitter (emission of electrons, $E_{\text{max}} = 293 \text{ keV}$), but there is a potential hazard because of its long half-life of 2.13·10$^5$ years, which brings severe problems for nuclear waste management. Therefore, a fundamental understanding of the chemistry of $^{\text{99}}\text{Tc}$ is essential to avoid its release into the environment. Furthermore, profound insights into fundamental and environmental chemistry of $^{\text{99}}\text{Tc}$ are important to gain a broad knowledge of potential mobilization mechanisms, distribution pathways and hazards after accidental release. The development of technetium
chemistry has been lagging behind that of its heavier congenor rhenium, primarily because the inherent radioactivity of all Tc isotopes has limited the number of laboratories that can study the chemistry of this fascinating element. Good examples for these curious gaps in fundamental technetium chemistry are the recent preparations of the simple trivalent technetium halides TcBr₃ and TcCl₃,[4,5] and the binary halide TcCl₆.[6]

Most current studies with technetium focus on the development of novel compounds for radiopharmaceutical diagnostics in nuclear medicine.[7–12] Radioactivity remains one of the most sensitive methods for non-invasive imaging, where the nuclear isomer of ⁹⁹mTc (⁹⁹mTc), plays a prominent role. An estimated 90% of all diagnostic procedures in nuclear medicine, carried out worldwide involve ⁹⁹mTc compounds.[13] Because of its nearly ideal nuclear properties and widespread availability at modest cost, ⁹⁹mTc is said to be the ‘workhorse’ for nuclear medicine. The availability of the isotope from generators, in combination with the short physical half-life time (6 h) and the emission of low energy γ-rays (140.5 keV) make ⁹⁹mTc a very convenient and practical isotope. Nevertheless, a site-specific conjugation of a metalloradionuclide like ⁹⁹mTc to a targeting vector in a manner in which the product retains the affinity of the parent biomolecule is challenging.

To address this issue, one can take advantage of the rich coordination chemistry of technetium. Coordination of radio metals to an appropriate and strong chelator, conjugated to a targeting biomolecule (peptides, oligonucleotides or small molecules) is the most frequently applied labeling strategy and known as the bifunctional chelator (BFC) approach[14] (Fig. 2).

The BFC approach conveniently enables screening of a large number of targeting molecules with one particular metal core for which the chelator is designed. In the field of technetium chemistry, Tc⁶⁺ and Tc⁴⁺ cores have been widely explored to prepare molecular imaging probes.[7–10]

In the future, fundamental chemistry of Tc and applied radiopharmacy have to be pushed forward due to their importance for society. Therein, synergy plays an important role. On one hand, knowledge gained by radiopharmaceutical developments can draw a detailed picture of Tc(VII) chemistry. This could answer questions of permanent interest since it could induce possibilities for the synthesis of new starting compounds. Moreover, it could answer the question of a biological activation after its accidental release into the environment.

Besides [Tc⁶⁺]: only a relatively small number of high-valent Tc-complexes are known. This can be rationalized by the lack of a suitable precursor complex for the synthesis of technetium VII complexes. [Tc⁶⁺] is a very prominent compound. It can be prepared by oxidizing Tc metal in a stream of oxygen at 400–600 °C. Chemically it behaves as if it was built of [TcO₄⁻]³⁻. [Tc⁶⁺] can therefore be heterolytically cleaved by strong ligands such as CH₃CN.[16] In contrast to the homologous binary metal oxide, [Re₂O₇]²⁻, which exhibits an infinite polymeric solid-state structure, [Tc₂O₇]³⁻ has a discrete molecular composition.[17] Due to its molecular character, [Tc₂O₇]³⁻ is volatile. This fact and the moisture sensitivity of [Tc₂O₇]³⁻ make this compound an inconvenient precursor for the synthesis of high-valent Tc complexes. As an alternative, pertechnetic acid (HTcO₄) was used. HTO₄ can be prepared in situ by the addition of strong Brønsted acids (e.g. H₂SO₄ or HCl) to a solution of [TeO₄]²⁻ in dry organic solvents. However, HTO₄ is volatile as well and thereby not a suitable precursor due to radioprotection restrictions. Besides the lack of a suitable precursor complex for high-valent Tc chemistry, a (water) stable metal core, such as the fac-{(TcOCl)₄}⁺ core for low-valent Tc chemistry, that could act as a building block is not available. Only compounds containing the fac-{(Tc(NAr)₃)}⁺ or fac-{(TcOCl)₄}⁺ core have so far been found to be sufficiently resistant to reduction and hydrolysis.

The reaction of the moisture-sensitive [TcO₂(OSiMe)₄] with 2,6-diisopropylphenylisocyanate or 2,6-dimethylphenylisocyanate leads to elimination of the CO₂ and the formation of tris-imido complexes [Tc(NAr₂)₂(OSiMe)₄] (Ar = 2,6-diisopropylphenyl or 2,6-dimethylphenyl). [19] [Tc(NAr₂)₂(OSiMe)₄] is the precursor complex for a variety of new Tc-VIII complexes of the general formula [Tc(NAr₂)₂Cl₂]⁻ (X = alkyl, I, η¹-Cp) (Scheme 1).[19,20]

Surprisingly, this interesting metal core was not a subject of further studies and no syntheses for corresponding ⁹⁹mTc complexes were developed. Water-stable fac-{(TcNAr₂)₂}⁺ complexes such as the [Tc(NAr₂)₂]⁺ could have interesting properties for potential radiopharmaceutical applications. However, most stable high-valent tech-
Our interest in containing the replacement reaction (Scheme bond formation between Tc complexes and a substrate, without the need of a ligand with 2.2(3+2)-Cycloadditions of This mirrors the spatial demand of terminal oxo ligands when the metal is in its highest oxidation state +vii. In this case, ligand field effects play a minor role and size of an iodide ion. 

Due to the lack of a suitable precursor complex for ligand replacement reactions only two synthetic pathways are available for the synthesis of complexes containing the fac-{TcO3}+ core: The direct activation of [TcO2]− and the oxidation of stable Tc complexes which contain the metal atom at a lower oxidation state. All known fac-{TcO3}+ complexes have been synthesized along these two routes. However, structurally characterized fac-{TcO3}+ complexes are relatively rare. Davison et al. reported the synthesis of [99mTcO3(HB(pz))3] through oxidation of [99mTcClO3(HB(pz))] with nitric acid or from [HTcO4]. In situ generated, volatile HTcO4 was the precursor of compounds with the general formula [TcO3(tacn-R)] (R = alkyl, iPr, tBu, Ph). [28] In this context we reported mixed anhydrides [TcO3(OCCOPh)] and [TcO3(OBF3)]. [29] Due to the fact that the activated [TcO3] species, such as [TcO3(OBF3)]−, are highly reactive and water sensitive all these reactions were done in organic (dry) solvents. Therefore, a synthesis from water remained unknown until 2006. [30] 

The small tridentate ligand 1,4,7-triazacyclonane and its derivatives form very stable mono oxo TcVII complexes but also due to the unique reactivities of -{TcO3}+ complexes, however, this reaction type allows for a new way of covalent bond formation between Tc complexes and a substrate, without the need of a ligand replacement reaction (Scheme 3).
3. A New Labeling Strategy for Radiopharmaceutical Applications

The knowledge about stability and reactivity gained by the fundamental $^{99m}$Tc experiments described previously make the $\text{fac-}[\text{TcO}_3]^{+}$ core very attractive as a new synthon for radiopharmaceutical applications. Water-stable $\text{fac-}[\text{99mTcO}_3]^{+}$ complexes would pave the way for a new labeling strategy based on (3+2)-cycloaddition strategies and can provide this research with a new impetus of innovation. In addition to functionalizations via cycloaddition, fundamental studies with $^{99m}$Tc evidence that the tripodal ligand can carry a second, bioactive function. This may present an excellent opportunity for the synthesis of novel bifunctional imaging agents. With this concept, directly transferred to $[^{99mTcO}_3]$-labeling procedures, highly versatile, bifunctional radiolabels would become accessible. Whereas today’s synthetic strategies for the development of bifunctionalized Tc complexes are based on ligand replacement reactions at the metal core (2+1, 3+1, 3+2, 4+1) and the HYNIC approaches, the $\text{fac-}[\text{TcO}_3]^{+}$ complexes can easily and efficiently be bifunctionalized in a ligand-based fashion (Scheme 4).

This special feature of the $\text{fac-}[\text{TcO}_3]^{+}$ complexes can be used as a creative route to prepare novel molecular imaging probes and can help to overcome limitations given by the coordination chemistry of technetium.

A crucial point for the development of this new labeling strategy was the synthesis of a $[^{99mTcO}_3]^{+}$ complex. In contrast to $^{99}$Tc chemistry, $^{99m}$Tc chemistry has to be done at high dilution: $[^{99mTcO}_3]$ is eluted with a physiological sodium chloride solution from the so-called $^{99m}$Tc generator ($^{99}$Mo immobilized on an aluminum oxide column) and is obtained in concentrations of $10^{-9}–10^{-10}$ mol/l. Organic solvents, toxic oxidizing agents and non-pertechnetate starting compounds, make the previously described $^{99}$Tc approaches unfeasible for the synthesis of $^{99m}$Tc trioxo analogues. For the synthesis of water-stable $^{99m}$Tc trioxo complexes ($[^{99mTcO}_3](\text{tacn-R}))^{+}$), a new reaction had to be developed starting from the diluted aqueous solution of $[^{99mTcO}_3]$ and using phosphines as reagents.[31] The phosphine was conjugated to a solid-phase support (resin), since the resin can easily be filtered after a heterogeneous reaction, which allows for the removal of excess reagents. In the presence of such polymers and a very strong triodal ligand such as tacn, the corresponding trioxo complex was formed in moderate yield after 1 h at 95 °C as the only product from $[^{99mTcO}_3]$ (Scheme 5).

The only suitable analytical method at these low concentrations is a HPLC system which is coupled to a UV- and a radio-detector. In general, the comparison of the HPLC retention times for the new $^{99m}$Tc compounds with a corresponding ‘cold’ standard compound (often the rhenium analog) confirms its identity. In this particular case the corresponding $^{99m}$Tc compound is the ‘cold’ standard (Fig. 5).

The synthesis of $[^{99mTcO}(\text{styrSO}_3)](\text{tacn})^{-}$ by the reaction of $[^{99mTcO}(\text{tacn})]^{-}$ with sodium 4-vinylbenzenesulfonate (Na(styrSO$_3$)) showed the suitability of

![Scheme 3. a) Cis-dihydroxylation reactions of alkenes with [OsO$_4$]; b) cis-dihydroxylation reactions of alkenes with [TcO$_3$(tacn)].](image)

![Scheme 4. Bifunctionalization of $[^{99mTcO}_3]^{+}$ by cycloaddition.](image)

![Scheme 5. Synthesis of $[^{99mTcO}_3]^{+}$-type complexes.](image)
In general, the (3+2)-cycloaddition of

The synthesis of $[^{99m}TcO_3(tacn)]^+$ by the reaction of $[^{99m}TcO_3(tacn)]^+$ with sodium 4-

Figure 5: HPLC traces (TEAP gradient) of a co-injection of $[^{99m}TcO_3(tacn)]^+$ and $[^{99TcO_3(tacn)}]^+$: the hypoxia-imaging agent,

This potential should be used to design new radiotherapeutics derived from this core. It is an important milestone for the development of potential radiopharmaceuticals since the biological profile of the new labeling strategy is currently under way.

4. Perspectives

The synthesis of new $[^{99m}Tc]$ compounds suitable for selective and site-specific targeting (e.g. cancer cells) for molecular imaging is the driving force and major objective of today’s research with this element. Despite some very promising and innovative new approaches, Tc chemistry is still looking for a new commercial application. This fact is starting to affect research efforts with technetium and the number of young academics working in this important field has dramatically decreased during the last years. New impulses are needed to recall the importance of Tc chemistry and radionuclide chemistry in general. Fundamental chemistry has the unique potential to create new compounds starting at the very beginning. This potential should be used to design compounds with novel properties, which can lead to the development of innovative strategies and applications. The chemistry of high-valent $fac$-$[^{99m}TcO_3(tacn)]^+$ complexes, including the new concept for the labeling of biomolecules with small $[^{99m}TcO_3(tacn)]^+$ complexes by way of a (3+2)-cycloaddition reaction. It is a potential for the development of new radiopharmaceuticals derived from this core.

Very recently the development of the new labeling method was brought to the next level. To characterize the biological profile of $[^{99m}TcO_3(tacn)]^+$ and to evaluate its suitability for preparing bioconjugates via cycloaddition reactions, a protocol for the synthesis of $[^{99m}TcO_3(tacn)]^+$ was developed. The first biodistribution studies of $[^{99m}TcO_3(tacn)]^+$ suggest the stability of this complex at neutral pH and ambient conditions. This new field demonstrates the importance of fundamental research and may have the potential to reactivate research efforts on this element in the future.

To develop the high-valent Tc chemistry with $fac$-$[^{99m}TcO_3(tacn)]^+$ compounds further, multifarious tasks have to be accomplished in the near future. For the synthesis of bifunctional imaging agents, the coordination chemistry of functionalized 1,4,7-triazacyclononane derivatives, which are suitable for coupling reactions, need to be elucidated. Alterations of the tacn ligand itself by extending the ring size from nine to more atoms may provide a very different reactivity pattern and can, in addition, be functionalized at a carbon atom. Furthermore, to make the (3+2)-cycloaddition competitive with today’s already established labeling strategies, reaction rates of the (3+2)-cycloadditions have to be optimized. The rate constants of these cis-dihydroxylation reactions are highly substrate dependent. Aiming at an understanding of a structure-reactivity correlation, systematic kinetic investigations are necessary. Even more important, a milestone for the convenience and routine applicability of this new labeling technique is the development of a more straightforward $[^{99m}TcO_3(tacn)]^+$ synthesis at neutral pH and ambient conditions. Thereby, the functional group tolerance...
of the [\textsuperscript{99m}TcO\textsubscript{3}(tacn-R)]\textsuperscript{+} synthesis will be increased and make this strategy even more attractive for the labeling of biomolecules.

Finally, it has to be emphasized that the high-valent chemistry of technetium is not only relevant from a radiopharmaceutical point of view, but also with respect to the significant amounts of long-lived \textsuperscript{99}Tc produced in the world's nuclear power plants. The possibility to bind \textit{fac-}[TcO\textsubscript{3}]\textsuperscript{+} compounds \textit{via} (3+2)-cycloaddition covalently to a substrate, such as surface-modified nanoparticles or resins, without a ligand replacement reaction could pave the way for new opportunities in the field of technetium immobilization. Therefore, the direct activation of [\textsuperscript{99m}TcO\textsubscript{3}] and its efficient transformation into water stable \textit{fac-}[TcO\textsubscript{3}]\textsuperscript{+} compounds is of great interest.

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