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# Rapid transformation of inorganic to organic and plant-available phosphorous in soils of a glacier forefield

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## Abstract

Chemical weathering of rocks or sediments is extremely important for the generation of soils, for the evolution of landscape, and as a main source of inorganic nutrients for plant growth and therefore for life. Temporal trends in weathering mechanisms, plant succession and nutrients availability in cold environments can be successfully studied in soil chronosequences along a glacier forefield. In the present paper, this was carried out in the pro-glacial area of Morteratsch. Different forms of phosphorous in the soil, stream and spring water chemistry were investigated. Apatite constitutes the main source of P, but it occurs only as a minor accessory mineral phase in the granitic/gneiss parent material. The identification of apatite was performed using SEM-EDX and cathodoluminescence. Water chemistry data indicated some calcite dissolution at the earliest phase of exposure, pyrite and – on older surfaces increasingly – feldspar weathering. Apatite also seemed to contribute to Ca which is leached from the soils. The concentrations of dissolved P in the stream and spring waters were, however, extremely low (only a few  $\mu\text{g P/l}$ ). In the topsoil, the

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total stocks of P showed a slight decrease with time. Losses were rather difficult to detect even though the water fluxes through the soils and discharges are relatively high. Soil organic P is almost identical to the EDTA-extractable fraction. In an 11.5 ky-old soil outside the glacier forefield the concentration and proportions of organic P, EDTA-extractable P and inorganic P forms did not differ that much from the oldest soils (max. 150 y) of the glacier forefield. In the bulk soil, about 78% of total P was transformed into an organic form (40% already after 150 years) and, in the fine earth, about 81% (40 – 70% after about 150 years of soil evolution). Thus the P transformation reactions are shown to be very rapid, occur predominantly in the early phase of soil formation, and could be best described by an exponential decay model.

Keywords: Phosphorous, glacier forefield, weathering, apatite, initial soils

## Introduction

Since approx. 1850 AD, most Alpine glaciers show phases of marked recessions, temporarily interrupted by smaller glacier advances (Furrer, 1991; Maisch, 1992; Holzhauser and Zumbühl, 1996). Locally, large deglaciated surfaces arose where vegetation development, i.e. primary plant succession, could start. The glacier forefield is in most cases defined as the area between the present-day glacier extent and the distinct moraines deposited in the 1850s. Soils in pro-glacial areas in the Alps are, consequently, young and were formed over a time-span of about 150 years (the time since the end of the 'Little Ice Age' in the 1850s; Fitze, 1982). As a consequence of warming, additional areas will become ice-free and subject to weathering and soil formation.

Alpine glacier forefields provide unique opportunities to study initial soil formation, nutrient availability and plant successions along the chronosequences and consequently the development of life. Biogeochemical weathering processes in such environments initiate the mobilisation of biolimiting nutrients (phosphorus, iron, potassium, calcium) and give rise to primary production (Föllmi et al., 2009). Plant succession is highly influenced by the underlying substrate, its weatherability and other site factors. Anderson et al. (2000) and Matthews and Whittaker (1987), for instance, showed that factor complexes like snow melt, exposure, moisture, terrain and time

control vegetation succession. However, only few studies deal with the substrate characteristics (parent material and soils) in pro-glacial areas and little is known about the rates of physical and chemical changes of soils in such areas and their consequences for plant growth. Several authors (Egli et al., 2003; Hosein et al., 2004; Föllmi et al., 2009; Mavris et al., 2010 etc.) suggested that glacially-derived material is subjected to enhanced chemical weathering, starting immediately after deposition in the pro-glacial zone and subsequently continuing for thousands of years after glacier retreat. The pro-glacial area is a potential zone of high geochemical reactivity due to the availability of freshly-ground reactive material (subglacially derived), often high water:rock ratios and contact times, high permeability and a constant supply of dilute waters (meltwater and rain/snowmelt) to percolate through the deposits (Föllmi et al., 2009).

The weathering of phosphorus-bearing minerals in rock substrata and the subsequent release of P to soils and the aqueous environment is considered as the principal mechanism by which P is transferred from the lithosphere to the biosphere (e.g., Filippelli, 2008; Föllmi et al., 2009). The release of directly or potentially bioavailable P to the biosphere represents a key factor in driving primary productivity rates. According to Föllmi et al. (2009), cycles of glaciation-deglaciation may reactivate transport-limited weathering regimes by replacing old and deeply-weathered soils with immature pedospheres, in which the dissolution of glacially-derived detrital P proceeds rapidly and in which bioavailable P is efficiently transferred into the biosphere.

Availabilities of nitrogen (N) and phosphorus (P) are important determinants of primary ecosystem succession. On areas that have only recently been deglaciated, inorganic forms of N and P predominate, which originate either from parent rock weathering and/or deposition (Welch et al., 2010).

The aim of our study was consequently to see how quickly inorganic P is transformed in a glacier forefield exhibiting young soils at an initial stage of development into other chemical forms such as organic or plant-available P and whether such transformation processes are bound to a loss (leaching) of phosphorus. For this purpose, weathering mechanisms had to be studied and, therefore, both soil and water analyses had to be taken into account.

## Investigation area

1 The investigation area is the glacier forefield of Morteratsch in the Upper Engadine (Swiss Alps).

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3 The backwall of the glacier is formed by high mountain peaks of the South Raetian Alps, such as  
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5 Piz Bernina (4049m asl.) and Piz Palü (3901m asl). The surrounding region of Morteratsch belongs  
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7 to the Lower Austroalpine Bernina Nappe that comprises mainly various plutonic rock types, like  
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9 granodiorites and alkali-granites, and subordinate syenites, diorites and gabbros (Büchi, 1994;  
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11 Spillmann, 1993). The border of the pro-glacial area is given by terminal moraines deposited in the  
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13 1850s during the 'Little Ice Age' (Fig. 1). The actual length of this pro-glacial area is approx. 2 km  
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15 and has an area of 1.8 km<sup>2</sup>. The pro-glacial area is in a valley that runs N-S. The altitude ranges of  
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17 the area of interest were between 1900 m asl to about 2050 m asl. The Swiss Glaciological  
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19 Commission of the Swiss Academy of Natural Sciences scnat (Gletscherberichte, 1881–2008)  
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21 recorded historical and recent glacier length fluctuations (Fig. 1; see also Maisch et al., 1993).

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23 Alpine glaciers have fluctuated during the last 10 – 12 ky near the terminal moraines formed in the  
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25 year 1850. This has been shown by many geomorphologic and climatic studies (cf. Patzelt, 1977;  
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27 Fitze, 1982; Gamper, 1985; Magny, 1992; Maisch, 1992;). Consequently, the exposure age or the  
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29 age of soils outside the glacier forefield increases very abruptly.

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31 The glacial till in the glacier forefield consists of granitic and gneissic material and was produced  
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33 through glacial transport within a small area of relatively homogeneous parent material. The  
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35 dominant soil units (some areas even do not have a soil) in the pro-glacial area are Haplic  
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37 Fluvisols, Endoskeletal, Skeletic or Lithic Leptosols, Humi-skeletal Leptosols (Fig. 1), including  
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39 some sites with Rankers that have a weak B horizon (IUSS working group, 2007), and, very rarely,  
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41 Dystric and Endogleyic Cambisols (endoskeletal). The young soils that are close to the glacier  
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43 showed almost no morphologic signs of chemical weathering and alteration products.

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45 Primary plant succession of the pro-glacial area with its predominantly silicatic parent material  
46  
47 starts with the pioneer plant communities *Oxyrietum digynae* and *Epilobietum fleischeri* and ends  
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49 with the larch/Swiss stone pine forest (*Larici-Pinetum cembrae*) after substantially more than 150  
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51 years (Burga et al., 2010). The first plants, i.e. *Epilobium fleischeri*, *Oxyria digyna*, *Linaria alpina*,  
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53 *Saxifraga aizoides*, *Rumex scutatus* appear about 7 years after deglaciation and reach greater  
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55 cover-abundance values (about 5 – 25%) after about 27 years (Burga et al., 2010). The first  
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1 species of the short-living *Oxyrietum digynae* appears approx. 10 years after deglaciation and  
2 disappears approx. 30 years later. The first shrubs of willows, green alder, dwarf-shrubs and small  
3 larch trees are found on areas that have been ice-free for about 12 to 15 years. The establishment  
4 of *Larici-Pinetum cembrae* takes place after about 77 years.

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7 Present-day climatic conditions for the Morteratsch site are approx. 0.5 °C mean annual air  
8 temperature and approx. 1000 – 1300 mm mean annual precipitation as calculated by using data  
9 from the nearby meteorological stations Samedan and Bernina.

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12 Over the period 2003–2007 mean daily discharge has been in the region of 4 – 5 m<sup>3</sup>/s, but the  
13 2003–2007 July mean daily discharge was 11.2 m<sup>3</sup>/s with a maximum of 85 m<sup>3</sup>/s on 9 July 2004  
14 (Stott et al., 2008). The Ova da Bernina is part of the catchment of the River Inn, which is the major  
15 Alpine tributary of the upper Danube. A further 4 km downstream at Pontresina, the Swiss Federal  
16 Office for the Environment has maintained a gauging station since 1955 (No. 2262 Berninabach–  
17 Pontresina, 46°29'06"N, 9°54'35"E) at an altitude of 1804 m asl which drains a catchment area of  
18 107 km<sup>2</sup>, approximately 47% of which is snow- and ice-covered in summer (Stott et al., 2008).  
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## 34 **Materials and methods**

### 35 *Soil and water sampling*

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38 Surface soil samples (topsoils, A or AO horizon; soil depth usually between 0 to 5/10 cm) were  
39 collected from the whole pro-glacial area (Fig. 1) forming a soil chronosequence ranging from 0 to  
40 150 y. Ten of these sites were close to the monitoring sites of Mavris et al. (2010). Additional sites  
41 were sampled close to the glacier (sites B1 – B4) as well as another one (site AP) outside the pro-  
42 glacial area on a surface having an age of about 11.5 ky (Egli et al., 2003).

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45 Around two kilograms of soil material were collected per sample. In order to yield reasonable  
46 results, large soil sampling volumes are needed for soils in alpine areas (Hitz et al., 2002). Soil  
47 bulk density was determined by a soil core sampler (or by excavated holes having a volume of  
48 about 500 – 2000ml that were backfilled with a measurable volume of quartz sand).  
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52 Four different types of water were considered: spring water and tributary water in the glacier  
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1 forefield (SW); and ice water (ICE), glacial stream water (GWa, GWz, BO1) and tributary water  
2 outside the glacier forefield (TW). The two streams outside the pro-glacial area (TW) were sampled  
3 to compare waters from young and older surfaces. The surface outside the pro-glacial area has a  
4 much older age (10 – 15 ky; Fitze, 1982) than the pro-glacial area (maximum age = 150 y).  
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7 Water was collected during the summer periods of 2008 – 2010 (June – September), 2 – 5 times  
8 per sampling season. All sampling bottles were acid washed (HNO<sub>3</sub> 1M) and rinsed with double  
9 distilled water to prevent contamination.  
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12 All water samples were either collected in PE- or glass (for  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$ ) vials filtered with 0.45  $\mu\text{m}$   
13 pre-cleaned millipore filters to eliminate suspended particles and stored in the dark at 4 °C.  
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### 21 *Chemical, physical and mineralogical analyses*

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23 Electric conductivity of the water, pH, HCO<sub>3</sub><sup>-</sup> and temperature were measured directly in the field.

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25 Water-pH was measured using a field pH-meter Metrohm E 604, conductivity using a  
26 microprocessor conductivity meter (LF320) having the standard conductivity cell (TetraCon 325)  
27 and HCO<sub>3</sub><sup>-</sup> was determined by titration (using HCl).  
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32 Ca, Mg, K, Na and Sr were analysed by an AAS (Perkin Elmer Atomic Absorption Spectrometer)  
33 using graphite furnace for the Sr measurements. The anions Cl<sup>-</sup>, NO<sub>3</sub><sup>-</sup>, SO<sub>4</sub><sup>2-</sup> and PO<sub>4</sub><sup>3-</sup> were  
34 analysed using an ion chromatograph Metrohm 733 IC Separation Centre. Silicon was measured  
35 photometrically. In addition, some samples were analysed for Al, Ba, Fe, P, Si, Sr using  
36 inductively-coupled plasma optical emission spectrometry (ICP-OES; PE 3300DV). Charged  
37 organic compounds were derived from the difference of the equivalent sum of cations and anions.  
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45 The measurements of  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  were performed using a Cavity Ring-Down Spectroscopy  
46 (CRDS) Picarro L1102i liquid analyser (manufactured by PICARRO). The device was connected to  
47 a LC PAL liquid auto-injector (CTC Analytics AG, Zwingen, Switzerland), equipped with a 1.2  $\mu\text{l}$   
48 syringe (model 26P/-mm/AS, 7701.2 N CTC manufactured by the Hamilton Company of Reno,  
49 Nevada, USA) for the injection of water samples into the machine. Samples and standards were  
50 injected into 32 x 11.6 mm screw neck 1.5 ml vials filled with 1 ml of water having  
51 PTFE/silicone/PTFE septums. The analysis scheme was based on two laboratory standards  
52 calibrated with Standard Light Antarctic Precipitation (SLAP) and Vienna Standard Mean Ocean  
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Water (VSMOW) used for interpolation and a control laboratory standard not included in the calibration. The regression between measurements and known  $\delta$  values for calibration standards was used to convert the measured absolute isotopic ratios to delta values.

$$\delta^{18}\text{O} = \left( \frac{\text{O}^{18}/\text{O}^{16}(\text{sample})}{\text{O}^{18}/\text{O}^{16}(\text{STD})} - 1 \right) \times 1000 \quad (1)$$

Total C and N contents of the soil and parent material were measured using a C/H/N analyser (Elementar Vario EL). Soil pH (in 0.01 M CaCl<sub>2</sub>) was determined on air-dried fine earth samples using a soil:solution ratio of 1:2.5.

The particle-size distribution was determined on some selected soil samples. After a pre-treatment of the samples with H<sub>2</sub>O<sub>2</sub> (3%), particle-size distribution of the soils was measured using a combined method consisting of sieving out the coarser particles (2000 – 32 $\mu$ m) and the measurement of the finer particles (< 32 $\mu$ m) by means of an X-ray sedimentometer (SediGraph 5100).

Element pools in the soil (Ca, Mg, Na, K, Si, Fe, Al, Ti, and Mn) were determined by a total dissolution method for the parental material (Table 1). Oven-dried samples were dissolved using a mixture of HF, HCl, HNO<sub>3</sub>, and H<sub>3</sub>BO<sub>3</sub> (Hossner, 1996) in a microwave oven and at a pressure of c. 25 bar in a closed system. Concentrations were measured using a Perkin Elmer Atomic Absorption Spectrometer (AAS). The NH<sub>4</sub>OAc-EDTA extractable fraction of phosphorous (plant available fraction) was measured according to Lakanen and Erviö (1971). Concentrations were determined photometrically at wavelength 692 nm using ammonium molybdate. Total P in the soils was determined using H<sub>2</sub>SO<sub>4</sub> after heating the samples to 550°C (Kuo, 1996) while the inorganic P fraction was determined on unheated samples. The organic fraction was determined as the difference between total and inorganic P.

Due to the low occurrence of apatite, XRD was not the appropriate technique with which to detect it. Therefore, scanning electron microscopy (SEM) coupled with an energy-dispersive spectroscopy (EDX) and cathodoluminescence (CL) was used as an alternative. CL measurements were carried out using a hot cathode CL microscope HC1-LM (Neuser et al., 1995). The system was operated at 14 kV accelerating voltage and a current density of about 10  $\mu$ A/mm<sup>2</sup>.



1 Luminescence images were taken on-line during CL operations using a Peltier cooled digital video-  
2 camera (KAPPA 961-1138 CF 20 DXC). CL spectra were recorded in the wavelength range 320 to  
3 900 nm using an Acton Research SP-2356 digital triple-grating spectrograph with a Princeton Pixi  
4 256B Spec 10 CCD detector that was attached to the CL microscope by a silica-glass fibre guide.  
5 CL spectra were measured under standardised conditions (wavelength calibration by a Hg-halogen  
6 lamp, spot width 30  $\mu\text{m}$ , measuring time 1 – 5 s). The measurements were carried out at the TU  
7 Bergakademie Freiberg, Germany.  
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10 The analysis with SEM-EDX was performed using a Dual Beam Quanta 200 3D FEI, with Dual  
11 BSD detector and W emitter operating at an accelerating voltage of 20 kV. Investigations were  
12 performed at the Institute for Building Materials (ETH Zurich, Switzerland).  
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### 23 *Statistics*

24 The soil and water datasets were checked for normal distribution using a Shapiro-Wilk test  
25 (SigmaPlot 11.0 (Systat Software Inc.); Jann, 2005). This procedure was applied using a two-tailed  
26 test for significance. The datasets usually did not show a normal distribution. Consequently, the  
27 Spearman rank correlation coefficient was taken.  
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## 39 **Results**

### 40 *Soils*

41 Some main soil characteristics are given in Table 2. In general, they all show only a (very) weak  
42 weathering stage and are characterised as sand to sandy loam. The total phosphorous  
43 concentrations were in the range of about 160 to 820 mg/kg and varied considerably along the  
44 chronosequence in the glacier forefield (Fig. 2). A temporal trend (accumulation or depletion) could  
45 not be observed. Also outside the pro-glacial and in a much older soil (about 11.5 ky), the P  
46 concentration was still in a similar range.  
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58 However, if the inorganic and organic P fractions are compared, temporal changes then seem to  
59 be obvious. Already within 150 years of soil evolution, a marked decrease of the inorganic P  
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fractions could be detected while the organic fraction seems to be mirror-inverted. After about 150 years, at least half of the inorganic P fraction in the fine earth had been transformed into an organic P form (Fig. 2). The EDTA-extractable fraction (which includes the exchangeable fraction, organically-bound species and species adsorbed on oxyhydroxides; cf. Miller et al., 1986; Keller and Védý, 1994) of phosphorous seemed to be almost identical to the organic fraction. Similarly to the organic fraction, the EDTA-extractable P increased with the age of the soils (Fig. 3): after about 150 years, also this fraction included about 50% of the total phosphorous.

An estimate of the stocks of the individual P fractions in the fine earth and bulk soil (including soil skeleton) was performed using the following equations:

Fine earth: 
$$P_{stock,FE} = P_i d_i \rho_i (1 - RM) \quad (2)$$

Soil skeleton: 
$$P_{stock,SK} = P_i d_i \rho_i RM \quad (3)$$

Bulk soil: 
$$P_{stock} = P_{stock,FE} + P_{stock,SK} \quad (4)$$

where  $P_{stock}$  denotes the P abundance ( $g/m^2$ ) in the fine earth (FE) or soil skeleton (SK),  $P_i$  the concentration of the fraction  $i$  ( $g/t$ ),  $d_i$  the soil depth (0.05m),  $\rho_i$  = soil density ( $t/m^3$ ) and  $RM$  the mass proportion of soil skeleton (diameter > 2mm).

The stock calculations were related to a unit depth of 5 cm (as most soils were very weakly developed and had an A, OA or O horizon of about that depth): for this purpose, also the soil skeleton fraction for each soil had to be measured.

The total P stocks in the fine earth were usually in the range of 5 – 25  $g/m^2$  to the depth of 5 cm depth (Fig. 4) and did not change over time. After about 100 years of soil evolution, the proportion of the organic P fraction sharply increased and reached after 10ky about 81% of the total fine earth mass. If the soil skeleton fraction is included in these calculations, then the P stocks usually varied in the range of 20 – 35  $g P/m^2$  to the depth of 5 cm depth. With time, around 78% of the total P was transformed into organic P (Fig. 4). The total P stocks (including the soil skeleton) in the topsoil seemed, however, to decrease slightly within the considered time span of 11.5 ky.

The chemical (Table 1) and also the mineralogical composition (Table 3) of the soil material typically reflected its granitic origin. Quartz, K-feldspar and Na-rich plagioclase were determined as the most abundant phases, followed by Fe-rich mica (biotite) and epidote. Using

1 cathodoluminescence, apatite could be detected – in the parent material or very young soils and in  
2 the older soils in the glacier forefield (Fig. 5). Apatite, characterised by euhedral crystals with  
3 mostly a size of about 30  $\mu\text{m}$  (but in some cases up to almost 1 mm), sporadically occurred in the  
4 glacial till. Apatite had a strong yellow luminescence emission (Fig. 5). The diadochic substitutions  
5 within the apatite clearly reflected the geology of the granitoid bedrock (see, e.g. Zhang et al.,  
6 1985; Kempe and Götze, 2002). The investigated samples did not feature any rim (or growing)  
7 structures in the crystal, which is typical for a homogeneous composition of apatite.  
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### 16 *Water analyses*

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18 The overall mineralisation of the waters was generally rather low. The total amount of dissolved  
19 compounds was in the range of  $5 \times 10^{-5}$  to  $1 \times 10^{-3}$  mol/l. The mineralisation degree and the  $\delta^{18}\text{O}$   
20 values confirmed that the water originated from the glacier-catchment and that the residence time  
21 in the subsurface is low. The dominant cation in the spring and stream waters (usually more than  
22 50% of the equivalents of the positively-charged ions) was  $\text{Ca}^{2+}$  (Fig. 6). The other cations ( $\text{Mg}^{2+}$ ,  
23  $\text{K}^+$ ,  $\text{Sr}^{2+}$ ) were measured in low concentrations while  $\text{Na}^+$  was slightly more abundant.  $\text{HCO}_3^-$  and  
24  $\text{SO}_4^{2-}$  dominated the anions.  $\text{Cl}^-$  and  $\text{NO}_3^-$  were present in concentrations  $< 0.4 \text{ mmol/dm}^3$ . The  
25 lowest concentrations were, not surprisingly, measured in the ice-water (small stream on top of the  
26 glacier) that consists purely of ice and snow melting water.  
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39 In general, a (temporal) shift of the chemical composition was measured when correlated to the  
40 age of the surface. Ice water predominantly consisted of  $\text{Ca}^{2+}$ ,  $\text{HCO}_3^-$  and some  $\text{Cl}^-$  and  $\text{NO}_3^-$   
41 whereas the stream and spring-water in the glacier forefield and particularly the tributary waters  
42 outside the glacier forefield became enriched in  $\text{K}^+$  and  $\text{Na}^+$ . Consequently, the older the surface  
43 the higher is the relative abundance of  $\text{K}^+$  and  $\text{Na}^+$ . On very young surfaces, the water composition  
44 reflected more calcite (or epidote) weathering and on older surfaces more plagioclase, alkali  
45 feldspar and/or sheet silicates weathering. The increasing importance of silicate weathering with  
46 surface age is also documented in the dissolved Si concentrations that steadily increased with  
47 distance from the glacier (Fig. 7).  
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59 The phosphate concentration was always below the detection limit (100  $\mu\text{g/l}$ ) using Ion-  
60 Chromatography (the routinely-used method). Only a few micrograms of P were detectable in  
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1 water using ICP-OES (Fig. 8). There may well even be a tendency of decreasing P concentrations  
2 with increasing distance from the glacier snout; however, the number of observations was too low  
3 for establishing a statistically-relevant trend.  
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### 7 *Hydrology*

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9 The  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  are compared in Figure 9. All values are near the 'Global Meteoric Water Line'  
10 (GMWL; Aggarwal et al., 2005) and the 'Local Meteoric Water Line' (LMWL). A better correlation  
11 is, however, achieved using the LMWL. For northern Switzerland a LMWL was derived with the  
12 equation  $\delta^2\text{H} = 7.55 \times \delta^{18}\text{O} + 4.8$  (Pearson, 1991). Compared to the GMWL, the LMWL and  
13 consequently the measured samples indicated a more continental climate. The  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$   
14 values are, however, also influenced by seasonal variations (cf. Hindshaw et al., 2011), because  
15 the melting of ice and snow influences the water residence time and the water composition. The  
16 ice-water (melting water from the glacier surface), collected close to the glacier snout, usually had  
17 distinctly less negative  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  values than the glacial stream. This particular ice-water had a  
18 very local origin. With decreasing altitude, the  $\delta^{18}\text{O}$  becomes less negative (Pearson, 1991).  
19 Compared to the ice-water that derived from locally melting water (collected on top of the glacier  
20 front), the more negative values of the glacial stream consequently mean that a significant part of  
21 the waters derives from higher altitudes. The glacier front was at an altitude of about 2050 m asl  
22 and the hydrologic catchment reaches altitudes of up to 4000 m asl. The oxygen isotopes were  
23 measured twice in 2010 (August and September). For both sampling campaigns, the  $\delta^{18}\text{O}$  became  
24 less negative as a function of distance from the glacier front. The slight increase in the  $\delta^{18}\text{O}$  ratio in  
25 the pro-glacial area consequently means that the stream water receives some contribution from  
26 tributary waters and/or precipitation. From the measured  $\delta^{18}\text{O}$  increase within the pro-glacial area  
27 ( $\sim 0.3\text{‰}$ ), this contribution can be estimated to have a maximum value of c. 20%. The change in  
28  $\delta^{18}\text{O}$  could be also due to weathering; however, the water dilution effect is most likely too high and  
29 overshadows potential weathering effects. A correlation analysis of  $\delta^{18}\text{O}$  with all the different  
30 solutes did not reveal any significant relationship. However, our  $\delta^{18}\text{O}$  dataset is small and further  
31 data would be necessary to provide conclusive findings.  
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## Discussion

Soil formation and weathering progress rapidly in the glacier forefield of Morteratsch (Mavris et al., 2012). Due to the increasing vegetation cover with age of surface exposure and higher plant productivity, more organic carbon is transferred into the soils as they develop (Table 2). This affects weathering and consequently also the transformation of the phosphorous forms.

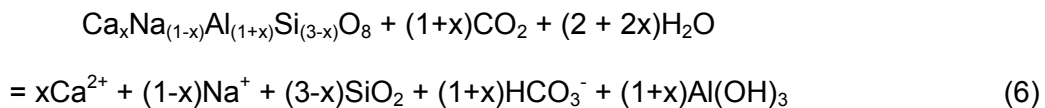
In granitic catchments, carbonate leaching is often seen as a major source of Ca (e.g. White et al., 1999a,b; Hosein et al., 2004). The investigated waters (springs, glacial stream, and rivers outside the pro-glacial area) are, however, strongly undersaturated with respect to calcite and dolomite (Mavris et al., 2010). Thus, the behaviour of  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  cannot be explained by only these two mineral phases. The solubility of Ca and Mg is also kinetically controlled by Ca- and Mg-bearing phases (silicates, phosphates). The dissolution of apatite represents a potential source of Ca (Blum et al., 2002). The correlation analysis showed that a certain relation – however not overwhelmingly significant – exists between  $\text{Ca}^{2+}$  and  $\text{PO}_4^{3-}$  (Table 4).

The runoff water reflects the geochemical processes in the glacier forefield. The main mineral dissolution reactions are (see also Mavris et al., 2010; Table 4):

Calcite (representing carbonates)

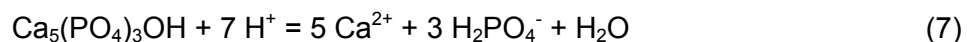


Plagioclase (representing tectosilicates):



where x is the Ca-rich component (White et al., 1999b).

Apatite (phosphate):



Elevated Ca/Si ratios e.g. Ca/ Si=12 (Anderson et al., 2000; Bench River) are thought to be caused by calcite dissolution. In general, calcite dissolution is a main supplier of Ca. However, this ultimately also depends on the amount of calcite present in the parent material. The bedrock or parent material contained very little calcite (Table 3). The Ca/Si ratios in the water usually were in

1 the range of 0.7 – 7 and consequently higher than in the Damma glacier forefield (Hindshaw et al.,  
2 2011). Therefore, some calcite dissolution is likely, but other Ca-bearing (unstable) minerals such  
3 as hornblende, plagioclase or apatite also have to be taken into account as possible Ca sources  
4 (Fig. 6). Sulphate significantly correlated with the major cations. Some pyrite (as a minor  
5 constituent of the parent material) armoured by calcite was also observed in the pro-glacial area.  
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7 The good correlation between  $\text{SO}_4^{2-}$  and  $\text{Ca}^{2+}$  most likely indicates the contemporaneous  
8 dissolution of pyrite and carbonates.  
9

10 In addition, biological processes contribute to element leaching from rocks and soils. The dissolved  
11 and charged organic compounds correlated with all major cations (Table 4). As reported by  
12 Brunner et al. (2011), pioneering fungi in a glacier forefield are able to dissolve the granite by  
13 releasing a variety of organic acids such as citric, malic and oxalic. In some cases, it was reported  
14 that mycorrhizal fungi produced oxalic acid that finally promote weathering (e.g. Hoffland et al.,  
15 2004; van Schöll et al., 2008). In particular Si and K mobilization might be affected by the presence  
16 of mycorrhiza (van Hees et al., 2006). Jongmans et al. (1997) showed that ectomycorrhizal fungi  
17 drill innumerable narrow cylindrical pores (diameter 3–10  $\mu\text{m}$ ) into weatherable minerals in podzol  
18 E horizons. The fungi probably form micropores by exuding strongly complexing low-molecular  
19 weight organic acids at their hyphal tips, causing highly local dissolution of silicates.  
20 Ectomycorrhizal (EM) fungi are known to colonize minerals in forest soil, and in many laboratory  
21 experiments it has been confirmed that EM fungi stimulate dissolution of minerals such as apatite,  
22 biotite and feldspars (Wallander and Hagerberg, 2004). Glacial deposits are at the beginning often  
23 depleted of symbiotic fungi but colonized by them over time after exposure (Cázares et al., 2005).  
24 Also weathering-associated bacteria produce oxalic acid or HCN (Frey et al., 2010). Frey et al.  
25 (2010) and Brunner et al. (2011) concluded that fungi and bacteria colonisation of young, granitic  
26 surfaces play a crucial role in the initial soil formation in previously glaciated mountains.  
27

28 Furthermore, one has to consider that the streamwater chemistry in glacier forefields may be  
29 strongly controlled by glacial melting and the routing of this glacial melt water through the glacier,  
30 resulting in marked seasonal and even diurnal variations (Hindshaw et al., 2011).  
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32 The progressive dissolution of apatite is the basic mechanism controlling the release of inorganic P  
33 in the newly-forming soil and also contributes to the availability of Ca for plants. Minor amounts of  
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1 P can probably be released by the decomposition of feldspar minerals. According to Smith (1983),  
2 feldspar minerals contain 0.00x-0.0x % P. Using cathodoluminescence, apatite could be detected  
3 (while this was not the case using XRD and DRIFT measurements) due to its bright yellow CL  
4 (Mn<sup>2+</sup> activation). The intensity of the luminescence is strongly orientation-dependent (Finch et al.,  
5 2003) and variation in luminescence can be linked to crystallographic controls (i.e. Barbarand and  
6 Pagel, 2001). In both samples, the parent material and the oldest soil of the glacier forefield, the  
7 presence of apatite could be verified.  
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10 Following dissolution from apatite, P is generally retained in Fe oxyhydroxides and organic matter  
11 (cf. Nezat et al., 2007). Also in this case, a significant correlation between the oxalate-extractable  
12 Fe and the EDTA-extractable P could be found ( $R^2 = 0.61$ ;  $p < 0.05$ ; Table 5). The Fe in biotite,  
13 when present, may also immediately lock up all of the P (Nezat et al., 2007). Föllmi et al. (2009)  
14 showed that with time organic- and iron-bound P almost completely replace detrital P. Apatite  
15 should be easily dissolved in surface waters due to its relatively high dissolution rate compared to  
16 silicate minerals. The solubility of P was, however, very low and the concentrations remained in the  
17 range of a few micrograms per litre. Also Hindshaw et al. (2011) measured very low (below the  
18 detection limit) P concentrations of  $< 50 \mu\text{g/l}$  in the stream waters in the Damma glacier catchment.  
19 As soils develop, primary apatite is weathered from the soil. The released Ca is taken up by  
20 vegetation, held on the exchange complex, or dissolved in groundwater and lost from the  
21 ecosystem. In contrast, P is comparatively immobile (Wood et al., 1984; Nezat et al., 2007) and,  
22 after being released by weathering, is retained in the soil complexed by organic matter and  
23 secondary minerals (Walker and Syers, 1976; Cross and Schlesinger, 1995), resulting in a  
24 relatively constant P concentration above the C horizon (Nezat et al., 2007). A rather weak  
25 correlation (error probability  $< 0.1$ ) of organic and EDTA-extractable P to pH exists whereas the  
26 org. C content obviously does not have a great effect on P speciation (Table 5). The decrease in  
27 pH (that is bound to calcite dissolution and the leaching of base cations) and the low effect of total  
28 organic C on P speciation shows that bacteria and fungi (including mycorrhiza), by the release of  
29 organic acids, are most likely decisive for the availability of phosphorous for plant growth  
30 (Jongmans et al., 1997; Frey et al., 2010).  
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1 The transformation of inorganic P forms into organic ones was a very rapid process and mainly  
2 occurred within 150 years of soil evolution in the glacier forefield of Morteratsch. The  
3 chronofunctions of the organic and EDTA-extractable P concentrations in the topsoil were best  
4 fitted to data using the exponential decay model (Fig. 3). The exponential decay model (cf. Lichter,  
5 1998) is given by  
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$$f(t) = a + (b - a)e^{-kt} \quad (8)$$

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11 where  $a$  represents an asymptote,  $b$  the initial quantity, and  $k$  the decay constant. Steady state  
12 conditions seem to be reached after about 100 – 200 years. A similar behaviour could be observed  
13 for the organic and inorganic P fraction (Fig. 4) where very rapid changes were measured and  
14 relatively soon an asymptotic end value seemed to be reached. The general behaviour and shape  
15 of the tendency seem to be similar to other components that have been measured in Alpine  
16 chronosequences (see Dahms et al., 2012).  
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25 Our data suggest that only a small part of P is lost due to leaching or weathering processes (Fig. 4)  
26 and that nearly all P cycles within the bio-geosphere – in contrast to many major ions that are  
27 leached with weathering. The decrease of the total P stocks (fine earth and soil skeleton included;  
28 Fig. 4C) in the topsoil within 150 years is however just slightly under the level of being statistically  
29 significant ( $p = 0.054$ ;  $R = -0.52$ ). Over 11500 years however, the error probability is slightly below  
30 5% ( $p = 0.03$ ;  $R = -0.56$ ) and the decrease consequently significant for total P in the topsoil. A part  
31 of this loss might be apparent, because with time bulk density and consequently the soil mass per  
32 unit area may change.  
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43 Phosphorous is very rapidly taken up by the plants, transferred back to the soil, retained and  
44 transformed (also microbiologically) into an organic form (Fig. 4B and D). A similar observation  
45 was also made by Tamburini et al. (2011) who showed that P is integrated in the soil by the  
46 microbial community; a process that begins very early in the soil development. It is also known  
47 from other investigations that organic matter accumulation is another P storage mechanism (see  
48 e.g., Bruland and Richardson, 2006). A low pH and a high content of Al and Fe oxihydroxides and  
49 organic matter were found to promote the phosphate sorption in European alpine soils (Kana et al.,  
50 2011). Our results, however, are in contrast to those of Eger et al. (2011) who measured rapid  
51 losses of total phosphorous at a rate of  $110 \text{ g m}^{-2} \text{ ky}^{-1}$  on sand dunes in a super-humid climate  
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(3455 mm annual precipitation) in the South Island of New Zealand. In this case, apatite was also the main P source. The rapid loss of apatite-P of over 50% after only 370 y of pedogenesis was almost completely offset by the rapid accumulation of organic P in the topsoil and subsoil. The inability of pedogenesis to sustain this transformation beyond 370 y of soil evolution resulted in a steep decline of P(tot). For Morteratsch, we can do the following calculations: assuming that the mean measured P-concentration in the waters (6µg/l) represents reality and that about 500 – 700 L water/m<sup>2</sup> percolate the soils each year, then a loss of 3 – 4.2 g m<sup>-2</sup> ky<sup>-1</sup> should be the result. Given the variability of total P in the soils (25 – 40 g P m<sup>-2</sup> per 5 cm depth), such a loss would be almost non-detectable even after 1000 years of soil evolution. However, after about 10 ky of soil evolution and using these calculations, the P-stock in the soil depth 0 – 5 cm should be close to zero. Long-term weathering rates of soils can also be derived using immobile element contents as indicators and the total content of the elements of interest (Table 6; Egli and Fitze, 2000). Ti was used as an immobile element. The parameter  $\tau_{j,w}$  represents the open-system mass transport function and is defined as:

$$\tau_{j,w} = \frac{C_{j,w}C_{i,p}}{C_{j,p}C_{i,w}} - 1 \quad (9)$$

where  $C_{j,p}$  (kg/t) is the concentration of element  $j$  in protolith (e.g., unweathered parent material, bedrock),  $C_{j,w}$  is the concentration of element  $j$  in the weathered product (kg/t),  $C_{i,p}$  (kg/t) the concentration of the immobile element  $i$  in the protolith (e.g., unweathered parent material, bedrock) and  $C_{i,w}$  the concentration of the immobile element  $i$  in the weathered product (kg/t). The open-system mass transport function calculates relative losses or gains. Absolute mass changes can be determined using the following equation:

$$\bar{m}_{j,flux}(z_w) = \sum_{a=1}^n C_{j,p} \rho_p \left( \frac{1}{\varepsilon_{i,w} + 1} \right) \tau_{j,w} \Delta z_w \quad (10)$$

$$\text{with } \varepsilon_{i,w} = \left( \frac{\rho_p C_{i,p}}{\rho_w C_{i,w}} \right) - 1$$

where  $\varepsilon_{i,w}$  represents the strain,  $\rho_p$  and  $\rho_w$  the bulk density ( $t/m^3$ ) of the protolith and the weathered soil, respectively,  $n$  = soil layers and  $\Delta z$  to the weathered equivalent of the columnar height (m) of a representative elementary volume.

The relative losses and gains ( $\tau_{j,w}$ ) for the oldest soil (Albic Podsol; 11.5 ky) are given in Figure 10. In contrast to the above-shown mass calculations (Fig. 4A and C), this approach considers also volume changes of the soil profile during its evolution ( $\varepsilon_{i,w}$ ). Due to podsolisation, Al, Fe and also base cations are strongly leached. Consequently  $\tau_{j,w}$  is negative down to the C horizon. This is typical for alpine and (strongly) weathered soils (see e.g. Egli et al., 2001). Phosphorous, however, has strongly positive values, except in the E horizon (strongly negative values due to the podsolisation effect). Two different variants (having a different starting point: C or BC horizon) were calculated – but in both cases a similar result was obtained. Phosphorous was not removed from the soil system. Although  $\tau_{j,w}$  is partially strongly positive, the total added amount of P over the time range of 11.5 ky is rather low: 10 – 39  $g/m^2$  in the depth range of 0 – 25 cm and about 100 – 250  $g/m^2$  over the whole profile. This again shows that losses over a long period of time must be minimal. In our case, either some aeolian contributions or some inhomogeneities within the granitic parent material overshadowed a potential P loss. In contrast to sites showing substantial erosion effects (cf. Föllmi et al., 2009) and a subsequent loss in detrital P, sites at stable topographic positions did not evidence a measurable loss, even after 11.5 ky.

As phosphorous strongly recycles within the vegetation and given the fact that with time the roots penetrate to greater soil depths, some phosphorous will be transferred back by litter to the topsoil which may therefore mask a potential P-leaching from the soils. Furthermore, with time more Fe-oxyhydroxides that strongly retain phosphorous are formed in the soil (Nezat et al., 2007). Extrapolating roughly the hydrology of the New Zealand sites (Eger et al., 2011) to Morteratsch (by increasing the percolation rate through soils to about 2800  $l/m^2$ ), then the P loss rates would be around 17  $g m^{-2} ky^{-1}$  to a depth of 5 cm depth and, consequently, still almost one order of magnitude lower than that measured by Eger et al. (2011).

## Conclusions

1 Phosphorous, as an intrinsic component for the development of life, is very quickly transformed  
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3 from inorganic into organic and plant-available forms. Within a time span of about 150 years of soil  
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5 formation in a glacier forefield, almost half of the inorganic fraction is made available for plant  
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7 growth. Apatite as the main P source (but still as a minor mineralogical component of the soil and  
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9 parent material) was detectable in very young soils and also after 150 years of weathering.  
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11 Leaching and consequent losses of P by percolating water through the soils seem to be very low:  
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13 however, this is not the case for other major ions and nutrients such as Ca, Mg, K and Na.  
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## References

- 37  
38  
39 Aggarwal, P.K., Gat, J.R., Froehlich, K.F., 2005. *Isotopes in the Water Cycle. Past, Present and*  
40  
41 *Future of a Developing Science*. IAEA. Springer, Dordrecht.  
42  
43 Anderson, S.P., Drever, J.I., Frost, C.D., Holden, P., 2000. Chemical weathering in the foreland of  
44  
45 a retreating glacier. *Geochimica et Cosmochimica Acta* 64, 1173-1189.  
46  
47 Armbruster, T., Bonazzi, P., Akasaka, M., Bermanec, V., Chopin, C., Gieré R., Heuss-Assbichler,  
48  
49 S., Liebscher, A., Menchetti, S., Pan, Y., Pasero, M., 2006. Recommended nomenclature of  
50  
51 epidote-group minerals. *European Journal of Mineralogy* 18, 551-567.  
52  
53 Barbarand, J., Pagel, M., 2001. Cathodoluminescence study of apatite crystals. *American*  
54  
55 *Mineralogist* 86, 473-484.  
56  
57 Blum, J.D., Klaue, A., Nezat, C.A., Driscoll, C.T., Johnson, C.E., Siccama, T.G., Eagar, C., Fahey,  
58  
59 T.J., Likens, G.E., 2002. Mycorrhizal weathering of apatite as an important calcium source in  
60  
61  
62  
63  
64  
65

base-poor forest ecosystems. *Nature* 417, 729-731.

1 Büchi, H. 1994. Der variskische Magmatismus in der östlichen Bernina (Graubünden, Schweiz).  
2  
3 Schweizer Mineralogische und Petrographische Mitteilungen 74, 359-371.

4  
5 Burga, C.A., Krüsi, B., Egli, M., Wernli, M., Elsener, S., Ziefle, M., Mavris, C., 2010. Plant  
6  
7 succession and soil development on the foreland of the Morteratsch glacier (Pontresina,  
8  
9 Switzerland): Straight forward or chaotic? *Flora* 205, 561-576.

10  
11 Bruland, G.L., Richardson, C.J., 2006. An assessment of the phosphorous retention capacity of  
12  
13 wetlands in the Painter Creek Watershed, Minnesota, USA. *Water Air and Soil Pollution* 171,  
14  
15 169-184.

16  
17  
18 Brunner, I., Plötze, M., Rieder, S., Zumsteg, A., Furrer, G., Frey, B., 2011. Pioneering fungi  
19  
20 from the Damma glacier forefield in the Swiss Alps can promote granite weathering. *Geobiology*  
21  
22 9, 266-279.

23  
24  
25 Cázares, E., Trappe, J.M., Jumpponen, A., 2005. Mycorrhiza-plant colonization patterns on a  
26  
27 subalpine glacier forefront as a model system of primary succession. *Mycorrhiza* 15, 405-416.

28  
29 Cross, A.F., Schlesinger, W.H., 1995. A literature review and evaluation of the Hedley  
30  
31 fractionation: applications to the biogeochemical cycle of soil phosphorus in natural ecosystems.  
32  
33 *Geoderma* 64, 197-214.

34  
35  
36 Dahms, D., Favilli, F., Krebs, R., Egli, M., 2012. Soil weathering and accumulation rates of oxalate-  
37  
38 extractable phases from alpine chronosequences of up to 1 Ma. *Geomorphology* 151-152, 99-  
39  
40 113.

41  
42  
43 Eger, A., Almond, P.C., Condron, L.M., 2011. Pedogenesis, soil mass balance, phosphorous  
44  
45 dynamics and vegetation communities across a Holocene soil chronosequence in a super-  
46  
47 humid climate, South Westland, New Zealand. *Geoderma* 163, 185-196.

48  
49  
50 Egli, M., Mirabella, A., Fitze, P. 2001. Weathering and evolution of soils formed on granitic, glacial  
51  
52 deposits: results from chronosequences of Swiss alpine environments. *Catena* 45, 19-47.

53  
54  
55 Egli, M., Fitze, P., 2000. Formulation of pedologic mass balance based on immobile elements: a  
56  
57 revision. *Soil Science* 165, 437-443.

58  
59  
60 Egli, M., Mirabella, A., Fitze, P. 2003. Formation rates of smectites derived from two Holocene  
61  
62 chronosequences in the Swiss Alps. *Geoderma* 117, 81-98.

63  
64  
65

- Filippelli, G. M., 2008. The global phosphorus cycle: past, present, and future. *Elements* 4, 89–95.
- 1  
2  
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52  
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55  
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57  
58  
59  
60  
61  
62  
63  
64  
65
- Finch, A.A., Hole, D.E., Townsend, P.D., 2003. Orientation dependence of luminescence in plagioclase. *Physics and Chemistry of Minerals* 30, 373-381.
- Fitze, P.F. 1982. Zur Relativdatierung von Moränen aus der Sicht der Bodenentwicklung in den kristallinen Zentralalpen. *Catena* 9, 265-306.
- Föllmi, K.B., Hosein, R., Arn, K., Steinmann, P., 2009. Weathering and the mobility of phosphorus in the catchments and forefields of the Rhône and Oberaar glaciers, central Switzerland: Implications for the global phosphorus cycle on glacial-interglacial timescales. *Geochimica et Cosmochimica Acta* 73, 2252-2282.
- Frey, B., Rieder, S.R., Brunner, I., Plötze, M., Koetzsch, S., Lapanje, A., Brandl, H., Furrer, G. 2010. Weathering-Associated Bacteria from the Damma Glacier Forefield: Physiological Capabilities and Impact on Granite Dissolution. *Applied and Environmental Microbiology* 76, 4788–4796.
- Furrer, G. 1991. 25000 Jahre Gletschergeschichte. Neujahrsblatt der Naturforschenden Gesellschaft in Zürich, Orell Füssli, Zürich.
- Gamper, M., 1985. Morphochronologische Untersuchungen an Solifluktioniszungen, Moränen und Schwemmkegeln in den Schweizer Alpen. Schriftenreihe Physische Geographie, 17, Zürich.
- Hindshaw, R.S., Tipper, E.T., Reynolds, B.C., Lemarchand, E., Wiederhold, J.G., Magnusson, J., Bernasconi, S.M., Kretzschmar, R., Bourdon, B., 2011. Hydrological control of stream water chemistry in a glacial catchment (Damma Glacier, Switzerland). *Chemical Geology* 285, 215-230.
- Hitz, C., Egli, M., Fitze, P. 2002. Determination of the sampling volume for representative analysis of alpine soils. *Zeitschrift für Pflanzenernährung und Bodenkunde* 165, 326-331.
- Hoffland, E., Kuyper, T.W., Wallander, H., Plassard, C., Gorbushina, A.A., Haselwandter, K., Holmström, S., Landeweert, R., Lundström, U.S., Rosling, A., Sen, R., Smits, M.M., van Hees P.A.W, van Breemen, N., 2004. The role of fungi in weathering. *Frontiers in Ecology and Environment* 2, 258–264.

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65
- Holzhauser, H., Zumbühl, H., 1996. To the history of the Lower Grindelwald Glacier during the last 2800 years - palaeosols, fossil wood and historical pictorial records - new results. *Zeitschrift für Geomorphologie N.F. Suppl.* 104, 95-127.
- Hosein, R, Arn, K., Steinmann, P., Adatte, T., Föllmi, K.B., 2004. Carbonate and silicate weathering in two presently glaciated, crystalline catchments in the Swiss Alps. *Geochimica et Cosmochimica Acta* 68, 1021–1033.
- Hossner, C.R., 1996. Dissolution for total elemental analysis. In: Sparks, D.L. (Ed.), *Methods of Soil Analysis, Part 3, Chemical Methods*. Soil Science Society of America Inc. and American Society of Agronomy Inc., Madison, WI, pp. 49–64.
- IUSS Working Group WRB. 2007. World Reference Base for Soil Resources 2006, first update 2007. *World Soil Resources Reports No. 103*. FAO, Rome.
- Jann, B., 2005. *Einführung in die Statistik. 2., bearbeitete Auflage*. Oldenburg Wissenschaftsverlag, München, Germany.
- Jongmans, A.G., van Breemen, N., Lundström, U.S., van Hees, P.A.W., Finlay, R.D., Srinivasan, M., Unestam, T., Giesler, R., Melkerud, P.-A., Olsson, M., 1997. Rock-eating fungi. *Nature*, 389, 682–683.
- Kana, J., Kopacek, J., Camarero, L., Garcia-Pausas, J., 2011. Phosphate sorption characteristics of European alpine soils. *Soil Science Society of America Journal* 75, 862-870.
- Keller, C., Védy, J-C., 1994. Distribution of copper and cadmium fractions in two forest soils. *Journal of Environmental Quality* 23, 987-999.
- Kempe, U., Götze, J., 2002. Cathodoluminescence (CL) behaviour and crystal chemistry of apatite from rare-metal deposits. *Mineralogical Magazine* 66, 151-172.
- Kuo, S. 1996. Phosphorus. In: Sparks, D.L. (Ed.), *Methods of Soil Analysis, Part 3-Chemical Methods (SSSA Book Series 5)*. Soil Science Society of America, Madison, pp. 869-919.
- Lakanen, E., Erviö, R., 1971. A comparison of eight extractants for the determination of plant-available micronutrients in soil. *Acta Agralia Fennica* 123, 223-235.
- Lichter, J., 1998. Rates of weathering and chemical depletion in soils across a chronosequence of Lake Michigan sand dunes. *Geoderma* 85, 255-282.

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61  
62  
63  
64  
65
- Magny, M., 1992. Holocene lake-level fluctuations in Jura and the northern subalpine ranges, France. Regional pattern and climatic implications. *Boreas* 21, 319-334.
- Maisch, M. 1992. Die Gletscher Graubündens. Rekonstruktion und Auswertung der Gletscher und deren Veränderungen seit dem Hochstand von 1850 im Gebiet der östlichen Schweizer Alpen (Bündnerland und angrenzende Regionen). Teil A: Grundlagen-Analysen-Ergebnisse. Teil B: Verzeichnisse-Datenkataloge-Gletscherkarten. *Phys. Geographie* 33. Geographisches Institut der Universität Zürich.
- Maisch, M., Burga, C.A., Fitze, P. 1993. Lebendiges Gletschervorfeld. Führer und Begleitbuch zum Gletscherlehrpfad Morteratsch. Engadin-Press, Samedan, Switzerland.
- Matthews, J.A., Whittaker, R.J., 1987. Vegetation succession on the Storbreen glacier foreland, Jotunheimen, Norway: A review. *Arctic and Alpine Research* 19, 385-395.
- Mavris, C., Egli, M., Plötze, M., Blum, J., Mirabella, A., Giaccai, D., Haeberli, W., 2010. Initial stages of weathering and soil formation in the Morteratsch proglacial area (Upper Engadine, Switzerland). *Geoderma* 155, 359-371.
- Mavris, C., Götze, J., Plötze, M., Mirabella, A., Haeberli, W., Egli, M., 2012. Weathering and mineralogical evolution in a high Alpine soil chronosequence: a combined approach using SEM-EDX, cathodoluminescence (CL) and Nomarski DIC microscopy. Submitted.
- Miller, W.P., Martens, D.C., Zelazny, L.W., 1986. Effect of sequence in extraction of trace metals from soils. *Soil Science Society of America Journal* 50, 598-601.
- Neuser, R.D., Bruhn, F., Gätze, F., Haberman, J., Richter, D.K., 1995. Kathodolumineszenz: Methodik und Anwendung. *Zentralblatt für Geologie und Paläontologie Teil 1*, 287-306.
- Nezat, C.A., Blum, J.D., Yanai, R.D., Hamburg, S.P., 2007. A sequential extraction to determine the distribution of apatite in granitoid soil mineral pools with application to weathering at Hubbard Brook Experimental Forest, NH, USA. *Applied Geochemistry* 22, 2406-2421.
- Patzelt, G., 1977. Der zeitliche Ablauf und das Ausmass postglazialer Klimaschwankungen in den Alpen. In Frenzel, B. (ed.), *Dendrochronologie und postglaziale Klimaschwankungen in Europa*. Erdwiss. Forschung, 13, Wiesbaden, 248-259.
- Pearson, F.J., Balderer, W., Loosli, H.H., Lehmann, B.E., Matter, A., Peters, Tj., Schmassmann, H., Gautschi, A., 1991. *Applied Isotope Hydrogeology - a case study in northern Switzerland*.

Studies in Environmental Science 43, Elsevier, Amsterdam, The Netherlands.

- 1 Smith, J.V., 1983. Some chemical properties of feldspars. Mineralogical Society of America,  
2  
3 Reviews in Mineralogy 2, 281-296.  
4
- 5 Spillmann, P. 1993. Die Geologie des penninisch-ostalpinen Grenzbereichs im südlichen  
6  
7 Berninagebirge. Dissertation ETH No.10175, Zürich.  
8  
9
- 10 Stott, T., Nuttal, A.-M., Eden, N., Smith, K., Mawell, D., 2008. Suspended sediment dynamics in  
11  
12 the Morteratsch pro-glacial zone, Bernina Alps, Switzerland. Geografiska Annaler 90A, 299-313.  
13
- 14 Tamburini, F., Bernasconi, S.M., Pfahler, V., Bünemann, E., Frossard, E., 2011. Where life meets  
15  
16 rocks: Understanding P cycling during the early phases of soil formation. Goldschmidt  
17  
18 Conferences Abstract, Mineralogical Magazine, 1983.  
19
- 20 van Hees, P.A.W., Rosling, A., Lundström, U.S., Finlay, R.D., 2006. The biogeochemical impact of  
21  
22 ectomycorrhizal conifers on major soil elements (Al, Fe, K and Si). Geoderma 136, 364-377.  
23
- 24 van Schöll, L., Kuyper, T.W., Smits, M.M., Landeweert, R., Hoffland, E., van Breemen, N., 2008.  
25  
26 Rock-eating mycorrhizas: their role in plant nutrition and biogeochemical cycles. Plant and Soil  
27  
28 303, 35–47.  
29
- 30 Walker, T.W., Syers, J.K., 1976. The fate of phosphorous during pedogenesis. Geoderma 15, 1-  
31  
32 19.  
33
- 34 Wallander, H., Hagerberg, D., 2004. Do ectomycorrhizal fungi have a significant role in weathering  
35  
36 of minerals in forest soils? Symbiosis 37, 249-257.  
37
- 38 Welc M., Bünemann, Frossard E., Jansa J., 2010. Enzymatic activities in the rhizosphere of  
39  
40 different plants at a glacier forefield. 19th World Congress of Soil Science, Soil Solutions for a  
41  
42 Changing World, pp. 48-51.  
43
- 44 White, A.F., Blum, A.E., Bullen, T.D., Vivit, D.V., Schulz, M., Fitzpatrick, J., 1999a. The effect of  
45  
46 temperature on experimental and natural chemical weathering rates of granitoid rocks.  
47  
48 Geochimica et Cosmochimica Acta 63, 3277-3291.  
49
- 50 White, A.F., Bullen, T.D., Vivit, D.V., Schulz, M.S., Clow, D.W., 1999b. The role of disseminated  
51  
52 calcite in the chemical weathering of granitoid rocks. Geochimica et Cosmochimica Acta 63,  
53  
54 1939-1953.  
55
- 56 Wood, T., Bormann, F.H., Voigt, G.K., 1984. Phosphorous cycling in a northern hardwood forest:  
57  
58  
59  
60  
61  
62  
63  
64  
65



biological and chemical control. *Science* 223, 391–393.

Zhang, S., Wang, L., Yang, W., 1985. Use of REE analysis in apatite to distinguish petrological and mineralogical series of granitic rocks. *Geochimica* 1, 45-57.

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**Table 1**[Click here to download Table: Table1.doc](#)

Table 1. Total chemical analysis of the parent material with standard deviation (SD; n=7) in the pro-glacial area.

	Fine earth	SD	Skeleton	SD	Bulk soil	SD
SiO <sub>2</sub> (g/kg)	683.5	97.1	626	52.8	661.6	58.6
CaO (g/kg)	12.1	5.6	37.9	14.7	24.0	8.9
MgO (g/kg)	12.5	5.4	19.1	7.4	15.7	5.3
K <sub>2</sub> O (g/kg)	35.9	4.4	33.1	6.3	34.5	4.9
Na <sub>2</sub> O (g/kg)	35.0	3.1	38.0	7.4	36.1	4.9
Al <sub>2</sub> O <sub>3</sub> (g/kg)	131.1	16.0	156.6	21.8	144.6	12.2
Fe <sub>2</sub> O <sub>3</sub> (g/kg)	26.1	7.2	38.0	12.4	32.1	9.8
TiO <sub>2</sub> (g/kg)	11.7	4.3	14.8	5.1	13.3	4.0
MnO <sub>2</sub> (g/kg)	0.7	0.2	0.9	0.2	0.7	0.2
P <sub>2</sub> O <sub>5</sub> (g/kg)	1.0	0.4	0.9	0.5	0.9	0.5

Table 2

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Table 2. Some physical and chemical properties of the investigated soils.

Site/Soil unit (IUSS, 2006)	Age (yr)	Horizons	Depth (cm)	BD <sup>1)</sup> (g/cm <sup>3</sup> )	Skeleton (wt. - %)	Sand <sup>2)</sup> (g/kg)	Silt <sup>2)</sup> (g/kg)	Clay <sup>2)</sup> (g/kg)	pH (CaCl <sub>2</sub> )	C (g/kg)	Fe <sub>o</sub> <sup>3)</sup> (g/kg)	P <sub>EDTA</sub> (mg/kg)	P <sub>org</sub> (mg/kg)	P <sub>inorg</sub> (mg/kg)	P <sub>tot</sub> (mg/kg)
S1/Humi-Skeletal Leptosol	138	O	0-6	0.2	41	n.m.*	n.m.	n.m.	4.6	321.1	1.028	210	130	333	463
		A	6-9	1.6	50	777	184	39	4.8	6.18	0.635	n.m.	n.m.	n.m.	n.m.
		BC	9-14	1.6	53	830	158	12	4.7	3.69	0.880	n.m.	n.m.	n.m.	n.m.
		C	14-30	1.8	40	757	222	21	4.6	2.8	0.588	n.m.	n.m.	n.m.	311
S2/Humi-Skeletal Leptosol	128	A	0-10	1.4	64	754	204	42	4.85	13.76	1.129	416	514	204	717
		AC	10-40	1.5	68	695	272	33	4.9	4.79	0.975	n.m.	n.m.	n.m.	n.m.
S3/Humi-Skeletal Leptosol	108	A	0-5	1.1	54	667	265	68	5.1	28.1	0.901	224	209	258	466
		AC	5-15	1.7	70	677	281	42	4.5	3.89	1.169	n.m.	n.m.	n.m.	n.m.
S4/Humi-Skeletal Leptosol	98	A/AC	0-5	1.5	53	939	61	15	5.2	14.6	0.414	138	160	173	333
		C	5-30	1.6	70	931	57	12	5.2	1.76	0.347	n.m.	n.m.	n.m.	218
S5/Humi-Skeletal Leptosol	68	A1/A2	0-4	1.3	3	530	432	38	4.85	13.82	0.742	158	154	156	310
		C1	4-9	1.4	36	573	387	40	4.65	2.53	0.583	n.m.	n.m.	n.m.	n.m.
		C2	9-20	1.8	64	570	372	58	4.6	3.65	1.465	n.m.	n.m.	n.m.	n.m.
S6/Skeletal Leptosol	48	A	0-3	1.6	64	n.m.	n.m.	n.m.	4.8	37.47	0.49	0	3	157	160
		C	3-25	1.7	68	852	129	19	5	2.1	0.444	0	0	345	345
S7/Skeletal	48	A	0-4	1.4	26	n.m.	n.m.	n.m.	6.1	60.24	0.371	108	173	629	802

Leptosol		C1	4-11	1.6	37	823	146	31	5.2	2.82	0.442	n.m.	n.m.	n.m.	n.m.
		C2	11-34	2.0	67	747	211	42	5.1	3.65	0.616	n.m.	n.m.	n.m.	n.m.
S8/Skeletal	58	OA	0-12	1.0	63	n.m.	n.m.	n.m.	4.6	118.2	1.226	247	267	548	815
Leptosol		C	12-33	1.8	48	712	220	68	4.4	3.12	0.839	n.m.	n.m.	n.m.	880
S9/Humi-Skeletal	73	OA	0-10	1.2	59	785	175	40	4.60	47.7	0.755	169	267	355	623
Leptosol		C	10-36	1.7	58	832	133	35	4.65	3.27	0.741	n.m.	n.m.	n.m.	n.m.
S10/Humi-Skeletal	78	A1/A2	0-5	1.3	60	818	143	39	4.6	24.5	0.787	33	134	417	551
Leptosol		AC	10-25	1.7	84	733	219	48	4.8	4.33	1.334	n.m.	n.m.	n.m.	754
BO1/no soil	1	C	0-5	1.4	73	n.m.	n.m.	n.m.	7.35	n.m.	n.m.	27	0	333	333
B2/no soil	3	C	0-5	1.4	75	n.m.	n.m.	n.m.	7.15	1.9	n.m.	241	13	701	714
B3/Lithic	20	(A)C	0-5	1.4	76	n.m.	n.m.	n.m.	6.05	2.9	n.m.	128	0	629	629
Leptosol															
B4/Skeletal	35	(A)C	0-5	1.4	71	n.m.	n.m.	n.m.	4.95	n.m.	n.m.	24	104	430	535
Leptosol															
AP/Albic	11500	O	0-5	0.8	14	n.m.	n.m.	n.m.	3.3	176.4	0.691	522	551	124	675
Podzol		E	5-10	1.3	30	542	375	83	3.3	16.4	0.269	n.m.	n.m.	n.m.	43
		Bs1	15-25	1.5	55	786	184	30	4.4	8.9	1.404	n.m.	n.m.	n.m.	693
		Bs2	40-60	1.7	57	n.m.	n.m.	n.m.	4.7	1.7	0.824	n.m.	n.m.	n.m.	690
		BC	70-80	1.7	53	780	187	33	4.9	0.0	0.391	n.m.	n.m.	n.m.	338
		C	110-120	1.7	67	777	191	32	4.9	0.0	0.336	n.m.	n.m.	n.m.	130

\* n.m. = not measured

<sup>1)</sup>BD = bulk density

<sup>2)</sup>Grain sizes: sand (2000-63  $\mu\text{m}$ ), silt (63-2  $\mu\text{m}$ ) and clay (< 2  $\mu\text{m}$ )

<sup>3)</sup>Oxalate-extractable fraction

**Table 3**[Click here to download Table: Table3.doc](#)

Table 3. Mineralogical composition of the fine earth fraction (<2mm) of the parent material (t = 0) and the most developed soil horizon (A horizon of a Humi-Skeletal Leptosol (Ranker), t = 138 y) within the proglacial area.

Minerals	Parent material (t = 0) wt.-%	A horizon (t = 138 yr) wt.-%
Quartz	29.1 (±0.8)	30.1 (±0.8)
Albite	27.0 (±1.3)	27.8 (±1.3)
Microcline	11.8 (±0.8)	12.0 (±1.0)
Muscovite	9.4 (±0.7)	10.5 (±0.8)
Epidote	8.0 (±0.7)	6.2 (±0.7)
Hornblende	5.5 (±0.6)	3.9 (±0.5)
Chlorite	1.8 (±0.7)	1.7 (±0.6)
Illite	1.4 (±0.6)	1.9 (±0.6)
Biotite	1.0 (±0.4)	0.7 (±0.5)
Calcite	0.2 (±0.3)	0.6 (±0.3)

**Table 4**  
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Table 4. Correlation matrix (Spearman rank correlation coefficient) of the investigated compounds in sources and river waters. The first number designates the correlation coefficient R (e.g. Ca-Mg: 0.91), the second the error probability (e.g. Ca-Mg: <0.01) and the third the number of observations (e.g. Ca-Mg: 95).

	Mg <sup>2+</sup>	Na <sup>+</sup>	K <sup>+</sup>	Sr <sup>2+</sup>	H <sup>+</sup>	H <sub>4</sub> SiO <sub>4</sub>	Cl <sup>-</sup>	NO <sub>3</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	HCO <sub>3</sub> <sup>-</sup>	Org <sup>n-</sup>	PO <sub>4</sub> <sup>3-</sup>
Ca <sup>2+</sup>	0.91 < 0.01 95	0.50 < 0.01 95	0.77 < 0.01 95	0.666 < 0.01 61	-0.45 < 0.01 95	-0.19 0.47 17	0.28 < 0.01 95	0.26 0.01 95	0.63 < 0.01 95	0.52 < 0.01 95	0.57 < 0.01 95	0.60 0.08 9
Mg <sup>2+</sup>		0.66 < 0.01 95	0.77 < 0.01 95	0.729 < 0.01 61	-0.34 < 0.01 95	-0.15 0.55 17	0.29 < 0.01 95	0.30 < 0.01 95	0.70 < 0.01 95	0.42 < 0.01 95	0.54 < 0.01 95	0.55 0.11 9
Na <sup>+</sup>			0.47 < 0.01 95	0.57 < 0.01 61	-0.19 0.066 95	0.08 0.76 17	0.27 < 0.01 95	0.35 < 0.01 95	0.64 < 0.01 95	0.27 < 0.01 95	0.31 < 0.01 95	0.15 0.68 9
K <sup>+</sup>				0.66 < 0.01 61	-0.30 < 0.01 95	-0.01 0.95 17	0.32 < 0.01 95	0.14 0.17 95	0.71 < 0.01 95	0.36 < 0.01 95	0.43 < 0.01 95	0.17 0.64 9
Sr <sup>2+</sup>					0.18 0.18 61	-0.14 0.59 17	0.52 < 0.01 61	0.51 < 0.01 61	0.73 < 0.01 61	-0.01 0.95 61	0.37 < 0.01 61	-0.25 0.49 9
H <sup>+</sup>						0.33 0.19 17	-0.05 0.64 95	-0.03 0.78 95	-0.32 < 0.01 95	-0.92 < 0.01 95	0.03 0.74 95	-0.49 0.17 9
H <sub>4</sub> SiO <sub>4</sub>							0.29 0.25 17	-0.15 0.55 17	-0.01 0.95 17	-0.24 0.34 17	0.28 0.26 17	- - -
Cl <sup>-</sup>								0.38 < 0.01	0.37 < 0.01	0.07 0.49	0.13 0.21	-0.06 0.84

	95	95	95	95	9
NO <sub>3</sub> <sup>-</sup>		0.47	0.11	-0.12	0.03
		< 0.01	0.30	0.26	0.91
		95	95	95	9
SO <sub>4</sub> <sup>2-</sup>			0.38	0.16	-0.13
			< 0.01	0.12	0.71
			95	95	9
HCO <sub>3</sub> <sup>-</sup>				-0.14	0.46
				0.16	0.17
				95	9
Org <sup>n-</sup>					0.34
					0.36
					9

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Table 5

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Table 5. Correlation matrix (Spearman rank correlation coefficient) of the investigated compounds in the soils (fine earth; only pro-glacial area). The first number designates the correlation coefficient R, the second the error probability and the third the number of observations.

	P(org)	P(inorg)	Fe	Al	Si	pH	org. C
P(EDTA)	0.70 < 0.01 15	-0.13 0.63 15	0.78 < 0.01 11	0.31 0.34 11	-0.17 0.60 11	-0.33 0.22 15	0.12 0.68 13
P(org)		-0.32 0.24 15	0.39 0.22 1	0.19 0.56 11	-0.36 0.26 11	-0.44 0.09 15	0.43 0.14 13
P(inorg)			0.00 0.99 11	0.26 0.42 11	0.26 0.416 11	0.49 0.03 15	-0.25 0.39 13
Fe				0.22 0.20 34	0.26 0.144 34	-0.44 < 0.01 34	0.26 0.14 34
Al					0.35 0.04 34	-0.32 0.06 34	0.34 0.05 34
Si						-0.09 0.61 34	-0.30 0.09 34
pH							-0.20 0.24 36

**Table 6**[Click here to download Table: Table6.doc](#)

Table 6. Total content of some elements in the bulk soil (fine earth and stone fragments up to about a diameter of 200 mm included) of the 11.5 ky-old Podsol outside the pro-glacial area.

Horizon	Depth cm	Fe g/kg	Al g/kg	Na g/kg	K g/kg	P g/kg	Ti g/kg
O	0-5	8.3	40.6	11.3	37.4	0.596	2.48
E	5-10	11.0	57.8	19.3	55.9	0.069	2.80
Bs1	15-25	14.3	68.9	21.3	58.6	0.409	2.78
Bs2	40-60	13.5	66.7	21.2	55.9	0.401	2.76
BC	70-80	11.0	63.6	21.9	58.2	0.254	2.66
C	110-120	13.2	70.3	22.2	58.1	0.162	2.55

Figure 1  
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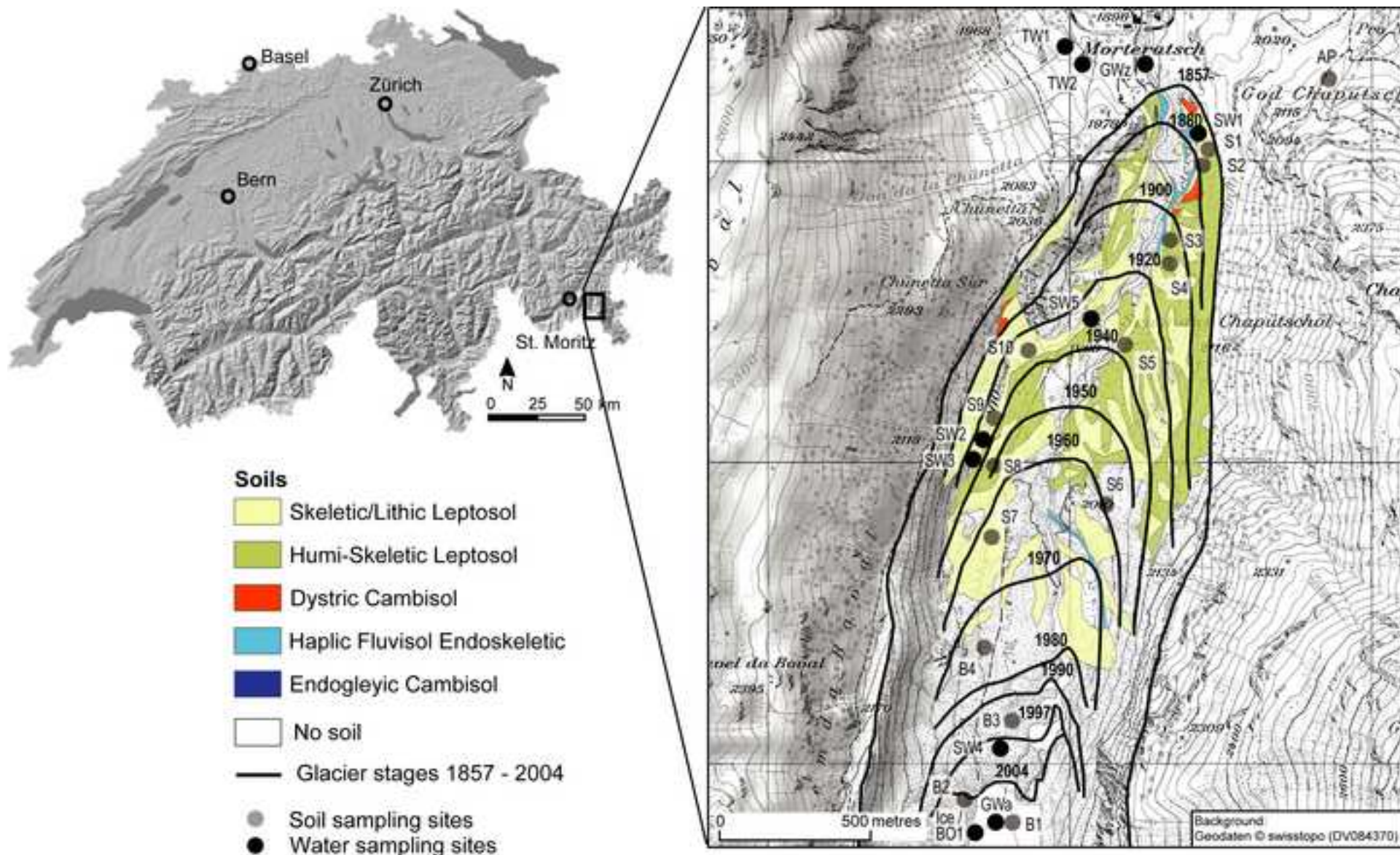
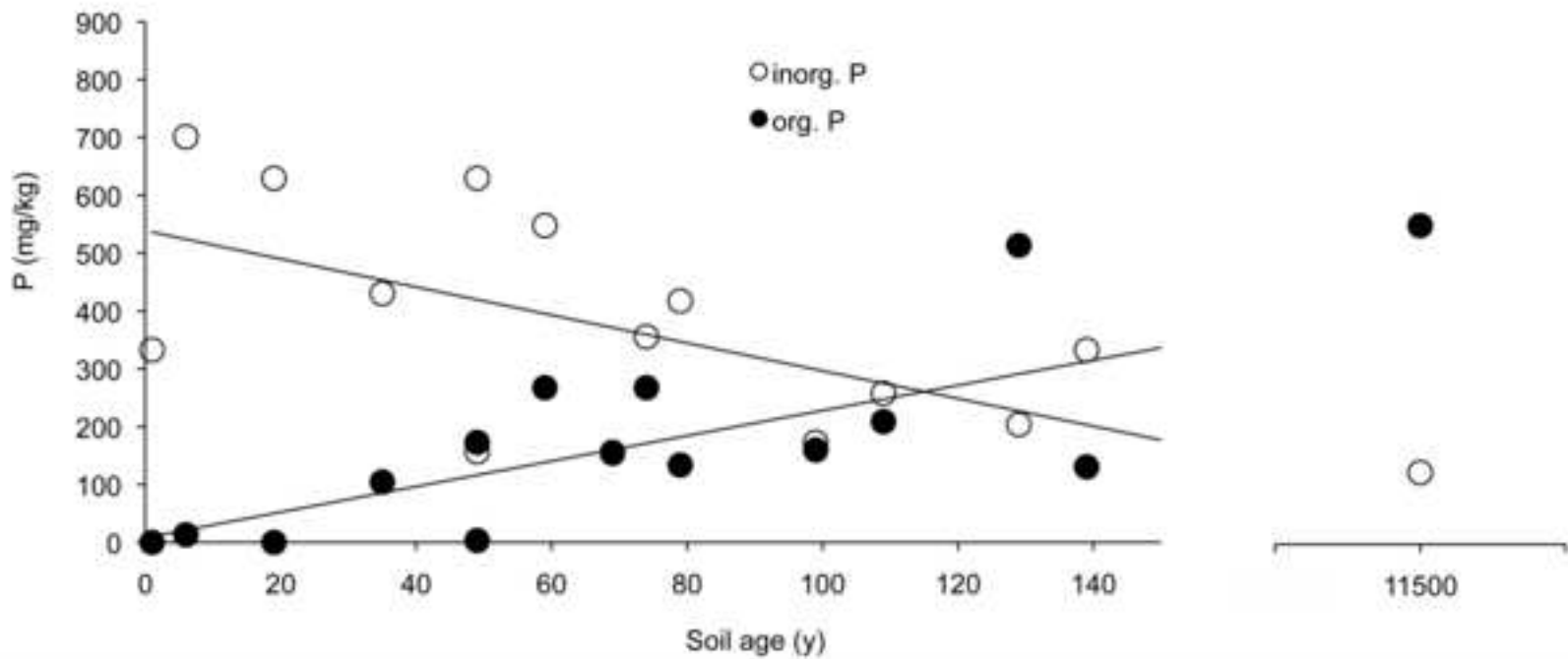
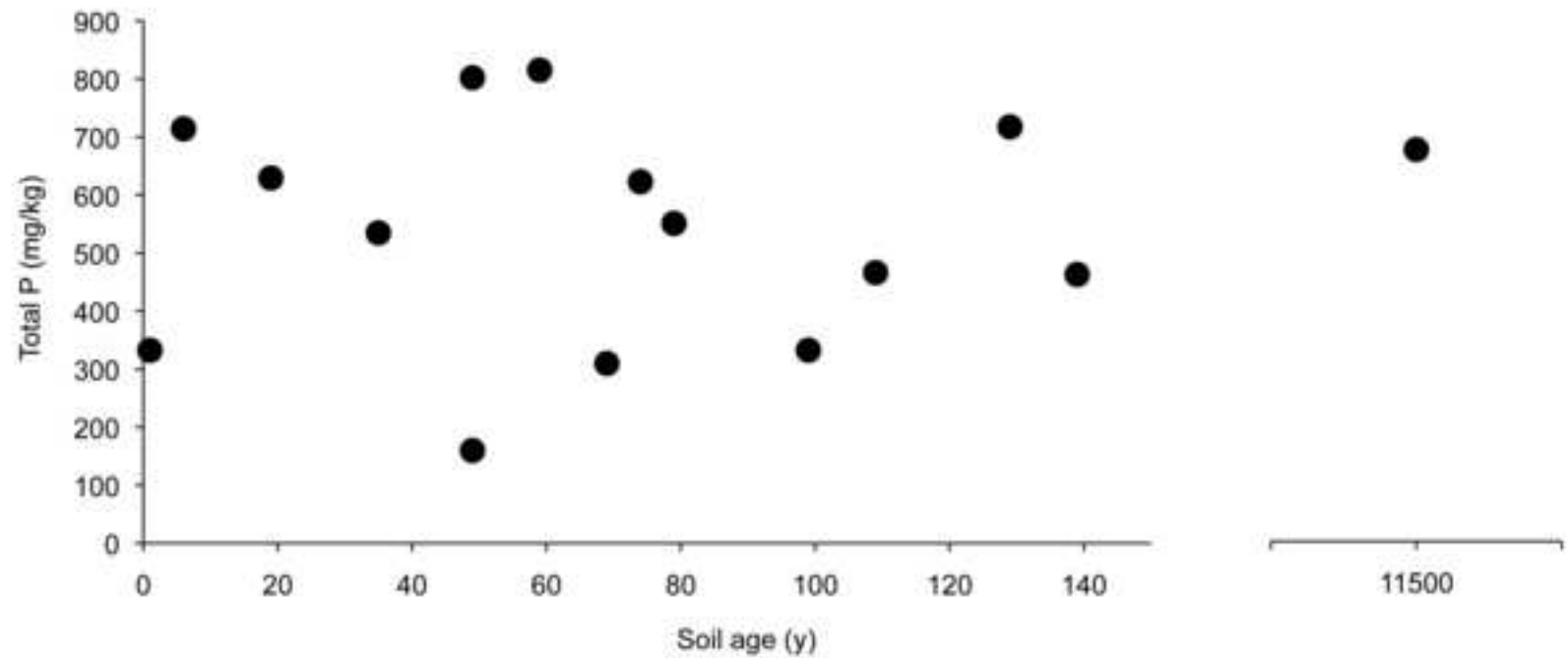
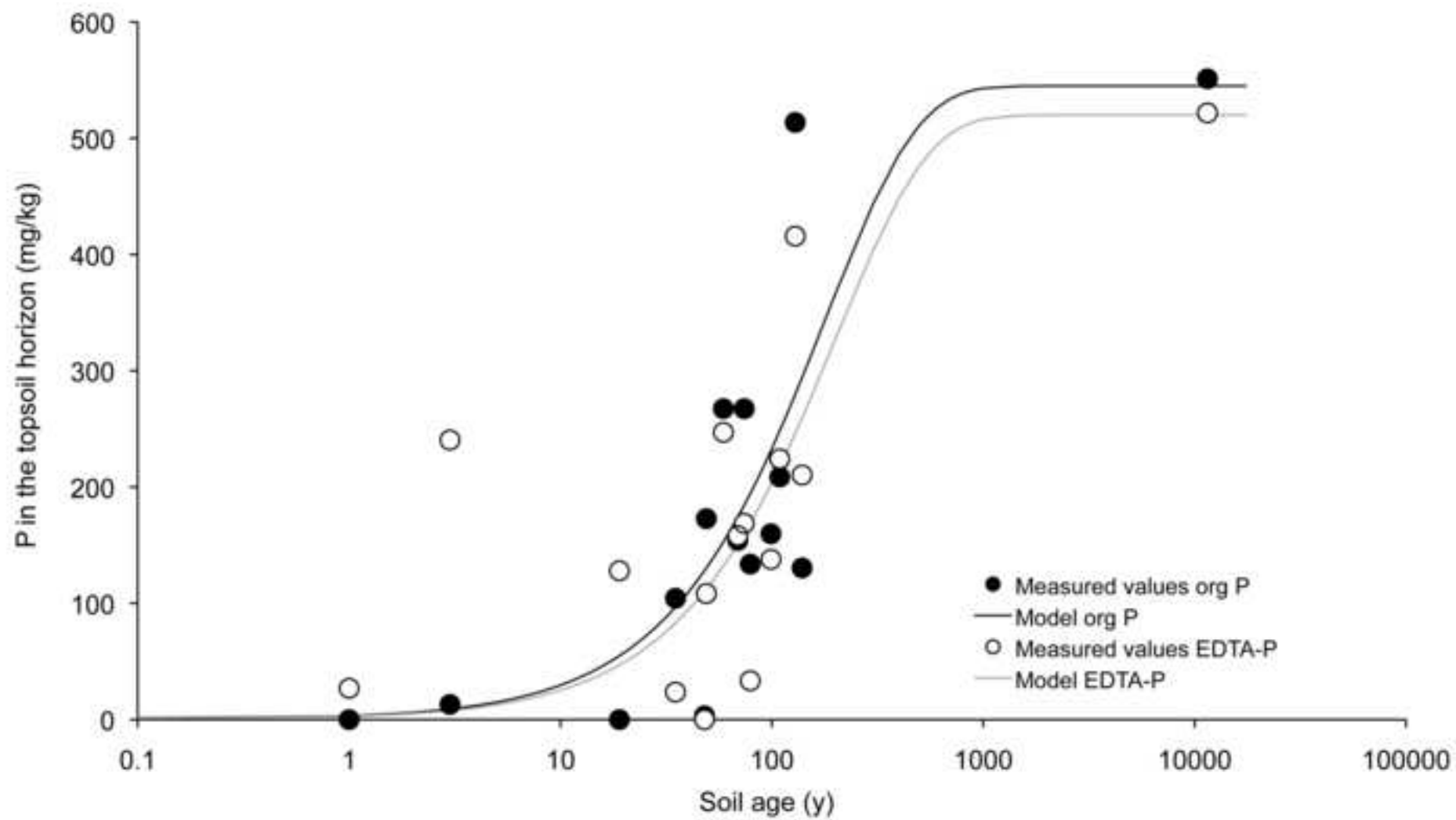


Figure 2  
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**Figure 3**  
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**Figure 4**  
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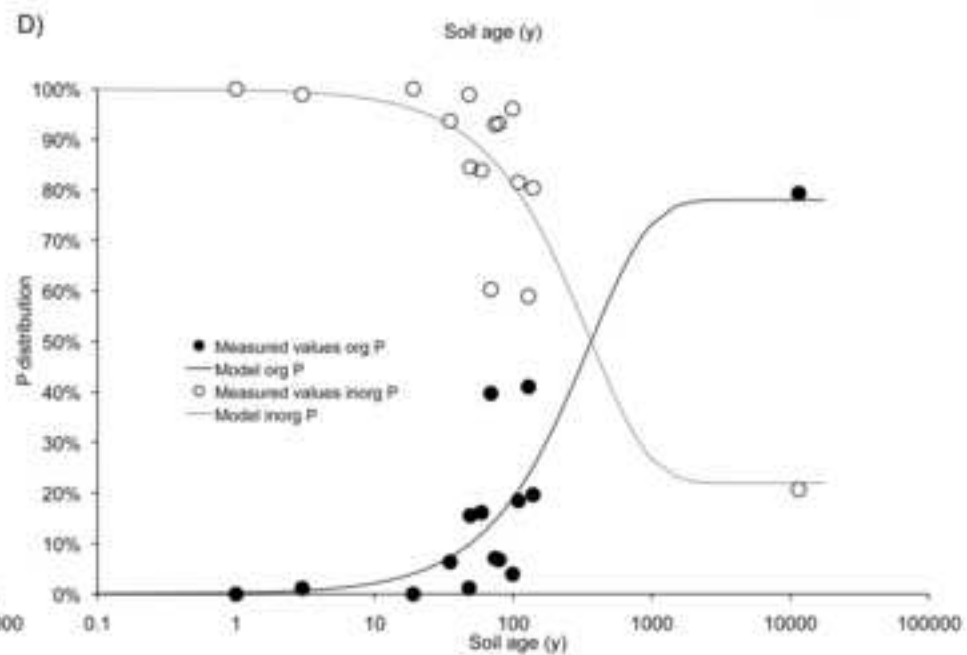
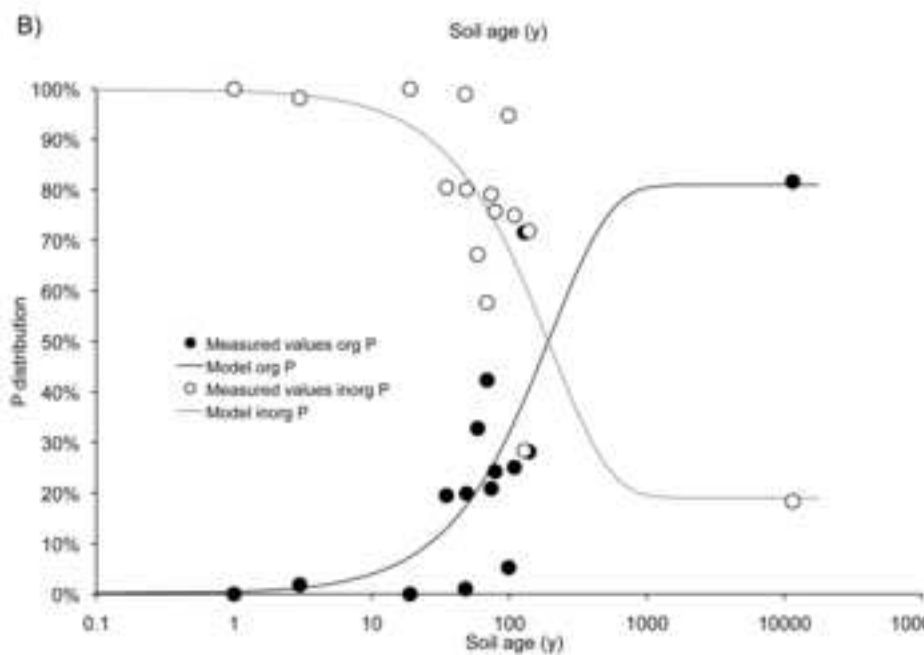
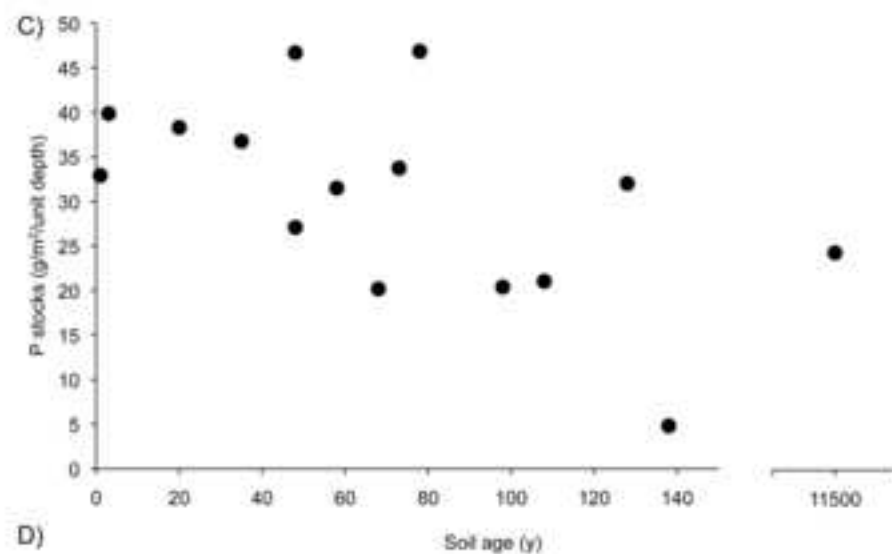
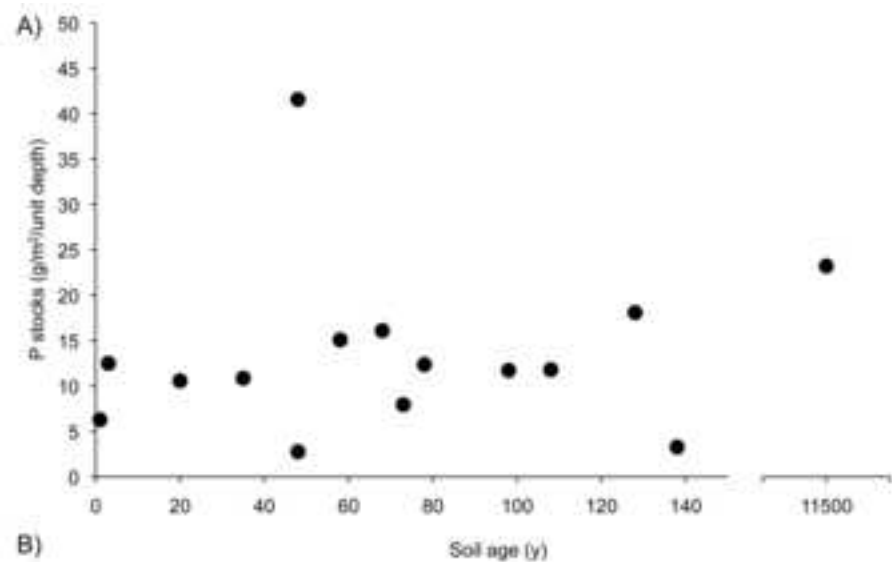




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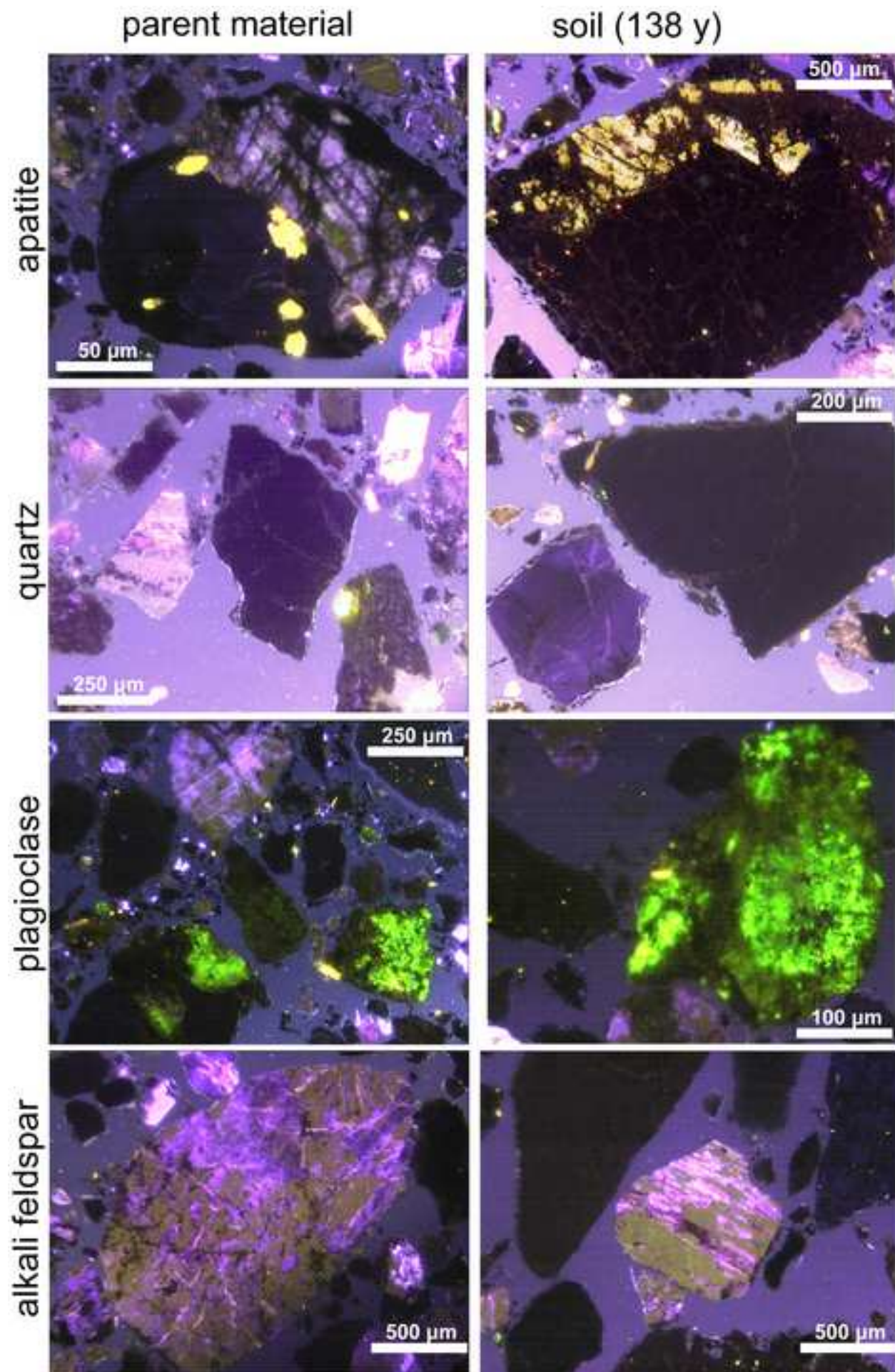


Figure 6  
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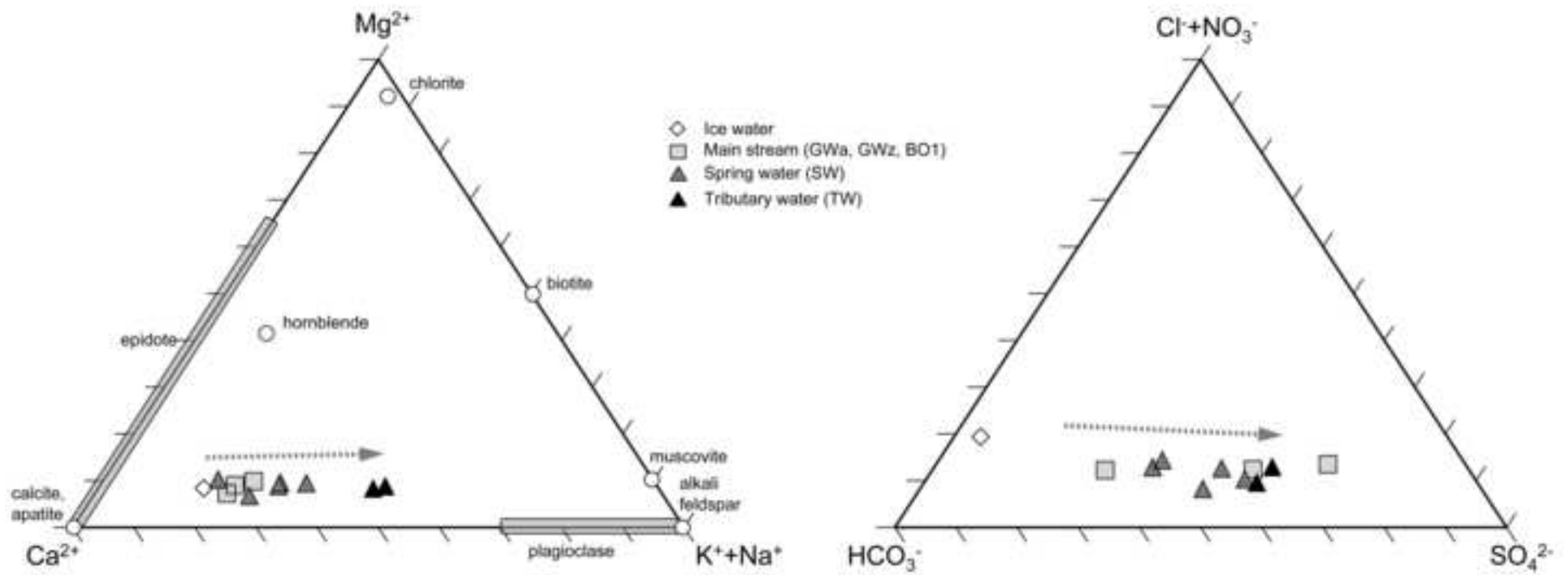




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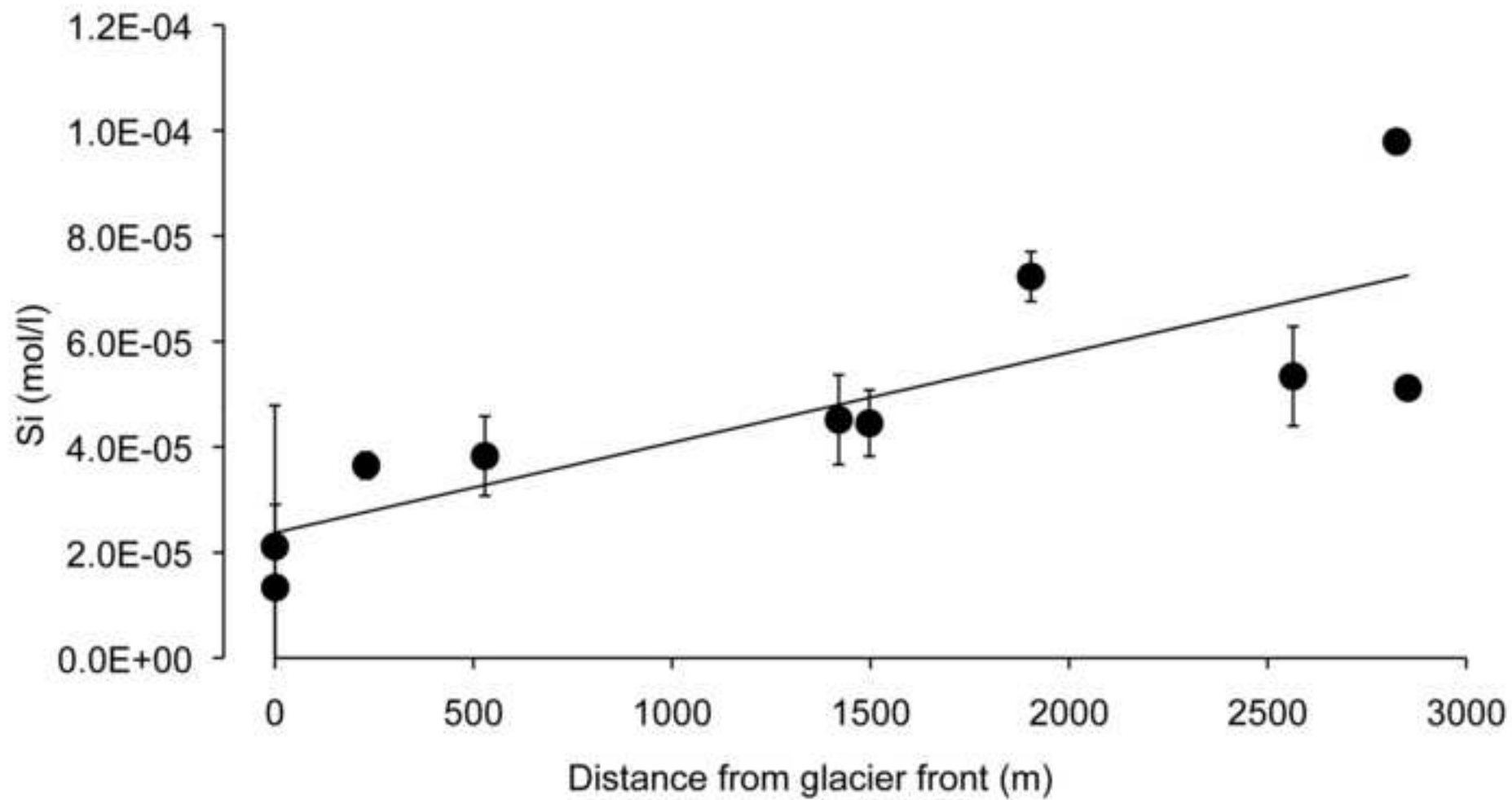


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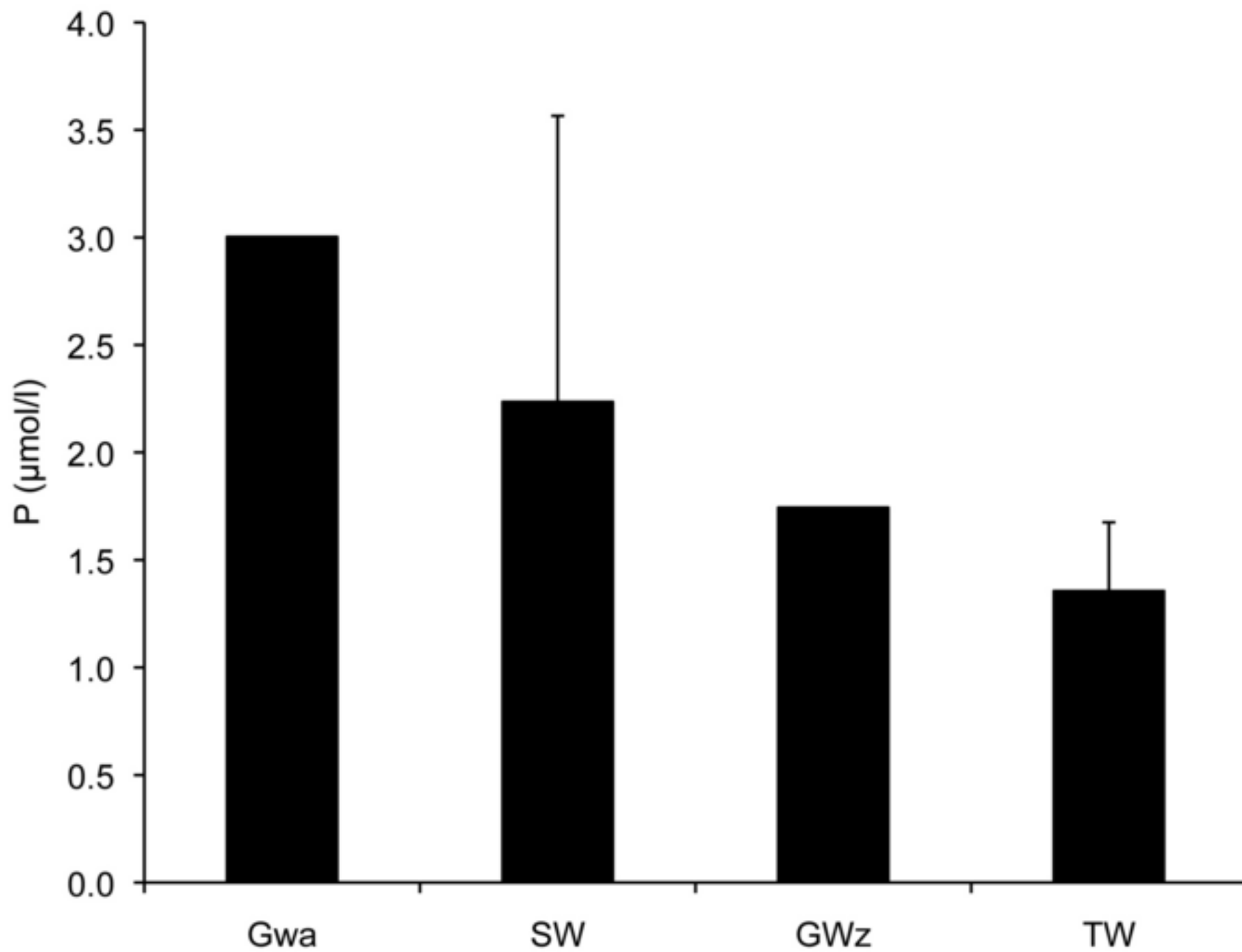


Figure 9

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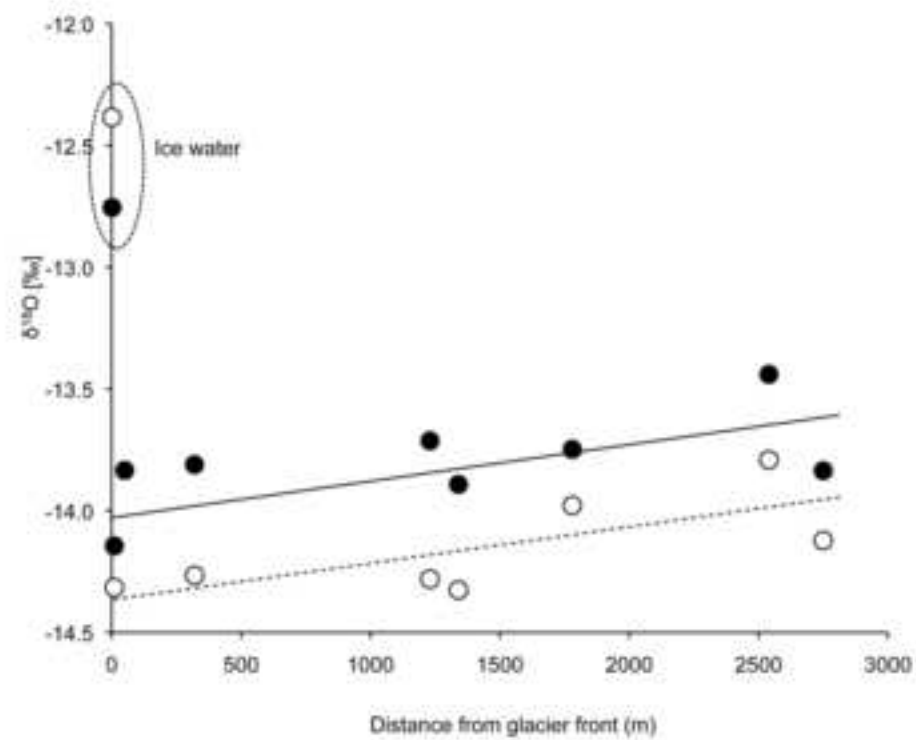
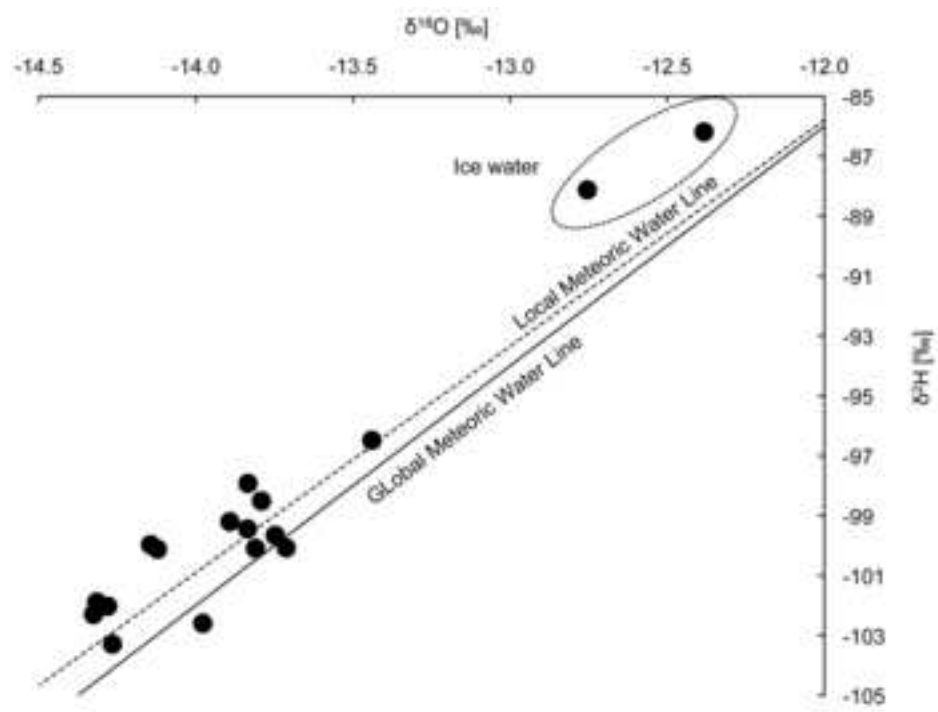
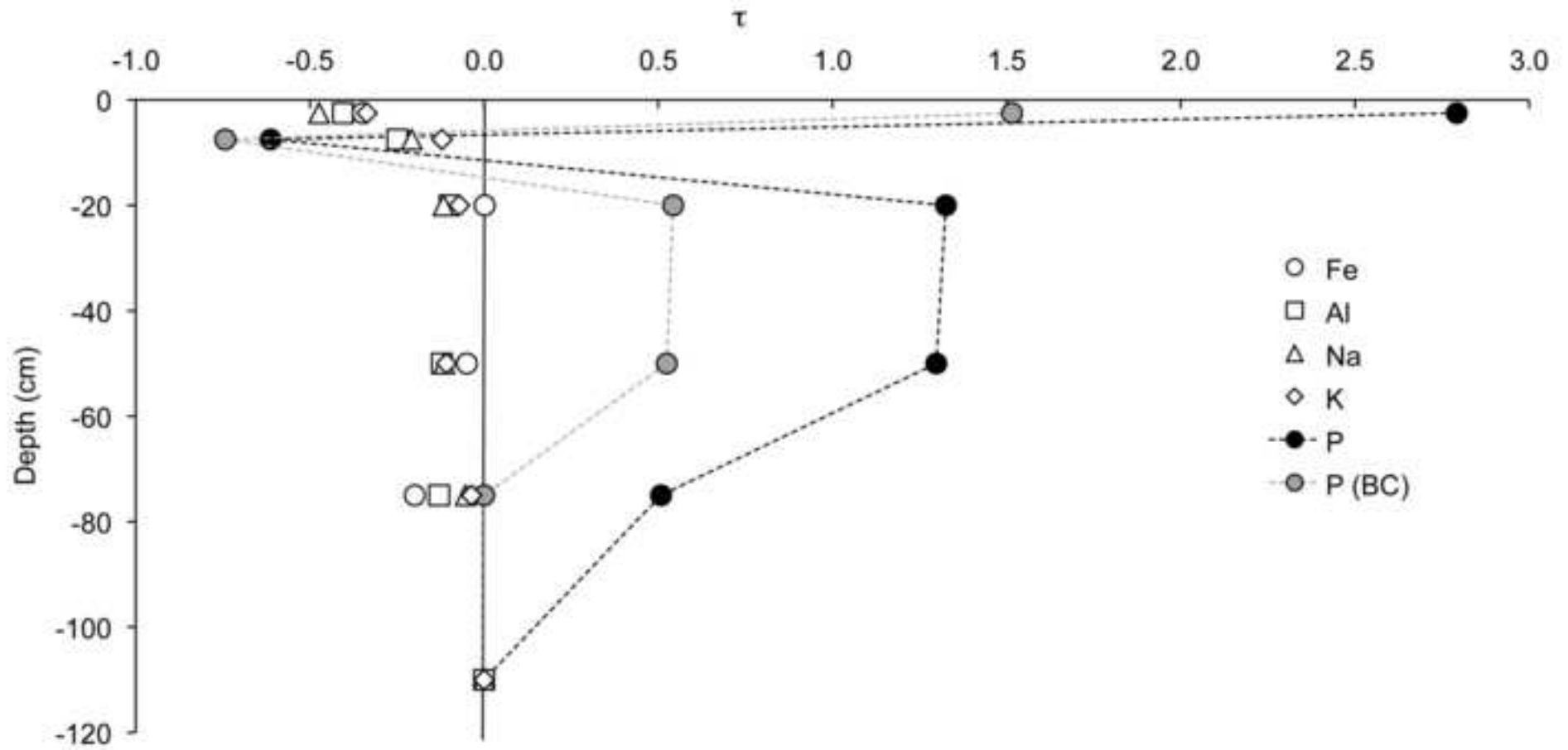


Figure 10  
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## Figure captions

Fig. 1. Location of the Morteratsch forefield with isochrones of glacier retreat, major soil units and position of the topsoil and water sampling sites. SW designates spring waters in the glacier forefield, TW are tributary waters outside the glacier forefield. The other water samples (GWA, GWz, BO1) were taken along the main river. Ice = streamlet on top of the glacier ice.

Fig. 2. Total, inorganic and organic P (fine earth) in the surface horizon of the soils as a function of time (with trend lines for the inorganic and organic P fractions).

Fig. 3. Relationship between EDTA extractable P and organic P. The concentration trend of organic and EDTA-P in the topsoil is modelled using an exponential decay model (see equation 8). The fitting parameters for org. P are:  $a = 545 \text{ mg/kg}$ ;  $b = 0.02 \text{ mg/kg}$ ;  $k = 0.00558 \text{ y}^{-1}$ , and for the EDTA-P:  $a = 520 \text{ mg/kg}$ ;  $b = 0.05 \text{ mg/kg}$ ;  $k = 0.00497 \text{ y}^{-1}$ .

Fig. 4. A) P-Stocks (unit depth: 0-5 cm) in the fine earth and B) relative distribution of the P fractions (inorganic and organic P fraction) in the fine earth. C) Total P stocks (including the skeleton fraction; unit depth: 0-5 cm) and D) relative distribution of the inorganic and organic P fraction in the bulk soil (including the soil skeleton). The trend of the relative distributions of the organic and inorganic fractions in B) and D) are modelled using an exponential decay model (see equation 8).

The fitting parameters in B) are:  $a = 81\%$ ;  $b = 0\%$ ;  $k = 0.0050 \text{ y}^{-1}$  for org. P and  $a = 19\%$ ;  $b = 100\%$ ;  $k = 0.005 \text{ y}^{-1}$  for inorg. P and in D)  $a = 78\%$ ;  $b = 0\%$ ;  $k = 0.0028 \text{ y}^{-1}$  for org. P and  $a = 22\%$ ;  $b = 100\%$ ;  $k = 0.0028 \text{ y}^{-1}$  for inorg. P.

Fig. 5. Cathodoluminescence micrographs of apatite, quartz, plagioclase and alkali feldspar of the parent material (0 y) and 138 y old topsoil. Apatite: yellow luminescing crystals due to the activation of  $\text{Mn}^{2+}$ ; quartz: blue-brown CL; plagioclase: green luminescing areas with strongly emitting  $\text{Mn}^{2+}$  peak;

alkali feldspar: violet emitting grains, related to both Al-O<sup>-</sup>-Al defects and Fe<sup>3+</sup> in Al-sites. The heterogeneous internal textures of the feldspar grains are caused by weathering of the material.

Fig. 6. Main chemical composition (mean values) of the analysed waters plotted in ternary diagrams (molar ratios). The dashed arrow shows the temporal evolution. The mineralogical compositions are from Nezat et al. (2007), except for epidote (Armbruster et al., 2006).

Fig. 7. Si concentrations (mean values and SD) in the water as a function of distance from the glacier front.

Fig. 8. Phosphorous concentrations in the waters: GWa = river water near the glacier front, SW = spring water in the glacier forefield, GWz = river water at the end of the glacier forefield, TW = tributary waters outside the glacier forefield.

Fig. 9. Relationship between  $\delta^{18}\text{O}$  and  $\delta^2\text{H}$  of the waters from the glacier forefield plotted together with the global and local water meteoric line.

Fig. 10. Open-system mass transport function  $\tau_{j,w}$  for Fe, Al, Na, K and P as a function of soil depth. Because the variability of the P concentration with soil depth is high, two different variants were calculated: one with the C horizon and another one with the BC horizon as the parent material (i.e. starting point).